

Standard Enthalpy of Formation (ΔH°_f) at 298 K

Formula	KJ/mol	Formula	KJ/mol	Formula	KJ/mol
Carbon		Chlorine		Sodium	
C (s, graphite)	0	Cl (g)	121.3	Na (s)	0
C (s, diamond)	1.88	Cl ₂ (g)	0	Na (g)	107.5
CH ₄ (g)	-74.6	HCl (g)	-92.3	NaCl (s)	-411.2
CH ₃ OH (l)	-238.6	Fluorine		Na ₂ CO ₃ (s)	-1130.7
C ₂ H ₂ (g)	227.4	F (g)	79.38	NaHCO ₃ (s)	-950.8
C ₂ H ₄ (g)	52.4	F ₂ (g)	0	Sulfur	
C ₂ H ₆ (g)	-84.68	HF (g)	-273.3	S ₈ (s, rhombic)	0
C ₂ H ₅ OH (l)	-277.6	Nitrogen		S ₈ (s, monoclinic)	0.3
C ₃ H ₈ (g)	-103.85	N ₂ (g)	0	SO ₂ (g)	-296.8
C ₃ H ₆ O (l, acetone)	-248.4	NH ₃ (g)	-45.9	SO ₃ (g)	-395.7
Hydrogen		NH ₄ (s)	-365.6	H ₂ SO ₄ (l)	-814.0

Standard Enthalpy Of Formation For Various Compounds

Lingjun Ying



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PHYSICAL CHEMISTRY R.C. SARASWAT, S.R. CHIKTE, M.P. SINGH, Unit I Thermodynamics I A Recapitulation of thermodynamic terms System surrounding types of system closed open isolated Thermodynamic variables intensive extensive properties thermodynamic processes isothermal adiabatic isobaric cyclic reversible irreversible State function path functions properties of state functions exact differential cyclic rule integrating factor concept of heat work 3L B Statements of first law of thermodynamics Definition of internal energy enthalpy heat capacity at constant volume at constant pressure Joule Thomson experiment Joule Thomson coefficient Inversion temperature calculations of W Q E H for expansion of gases for isothermal adiabatic conditions for reversible process carnot's cycle its efficiency thermodynamic scale of temperature 5L C Thermochemistry Heat of reaction standard states relation between heat of reaction at constant volume at constant pressure Hess's law of constant heat of summation its applications bond dissociation energy its calculations from thermochemical data variation of heat of reaction with temperature Kirchoff's equation 4L Unit II Thermodynamics II A Second law of thermodynamics Need for second law of thermodynamics statements of second law of thermodynamics concept of entropy entropy as a state function of V T P T entropy change in phase change for ideal gas entropy as criteria of spontaneity equilibrium 4L B Free energy functions Helmholtz free energy A Gibbs free energy G their properties standard free energies effect of temperature on free energy Gibbs Helmholtz equation its applications A G as criteria for thermodynamic equilibrium 4L C System of variable composition Partial molar quantities chemical potential Van't Hoff's reaction isotherm relation between standard free energy change equilibrium constant thermodynamic derivation of law of mass action effect of temperature on equilibrium constant reaction isochore 4L Unit III Phase Equilibria A Phase rule Statement of phase rule definition of phase component and degree of freedom derivation of phase rule Clapeyron equation its application in deciding slopes of line for two phase equilibria applications of phase rule to two phase equilibria of i water system ii sulphur system iii Pb Ag system 6L B Liquid Liquid mixtures Ideal liquid mixtures Raoult's law of ideal solutions Henry's law non ideal systems azeotropes HCl H₂O ethanol water system Partial miscible liquids Phenol water system trimethylamine water nicotine water system lower upper consolute temperature effect of impurity Immiscible liquids Steam distillation Nernst distribution law Limitations deviations applications 6L Unit IV Solid State Laws of crystallography i Law of constancy of interfacial angles ii Law of rationality of indices iii Law of symmetry symmetry of elements in crystals Unit cell space lattice orientation of lattice plane Miller indices Bravais lattices crystal systems X ray diffraction by crystal derivation of Bragg's equation

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