

17

THERMOCHEMISTRY

Reaction that releases heat is exothermic
 Reaction that absorbs heat is endothermic
 Exothermic: $\Delta H < 0$
 Endothermic: $\Delta H > 0$

Practice Problems

1. How much heat is released when 100.0 g of water is cooled from 25.0°C to 5.0°C?

SECTION 17.1 THE FLOW OF ENERGY—HEAT AND WORK

1. How much heat is released when 100.0 g of water is cooled from 25.0°C to 5.0°C?

2. How much heat is released when 100.0 g of water is cooled from 25.0°C to 5.0°C?

3. How much heat is released when 100.0 g of water is cooled from 25.0°C to 5.0°C?

4. How much heat is released when 100.0 g of water is cooled from 25.0°C to 5.0°C?

5. How much heat is released when 100.0 g of water is cooled from 25.0°C to 5.0°C?

6. How much heat is released when 100.0 g of water is cooled from 25.0°C to 5.0°C?

7. How much heat is released when 100.0 g of water is cooled from 25.0°C to 5.0°C?

8. How much heat is released when 100.0 g of water is cooled from 25.0°C to 5.0°C?

9. How much heat is released when 100.0 g of water is cooled from 25.0°C to 5.0°C?

10. How much heat is released when 100.0 g of water is cooled from 25.0°C to 5.0°C?

SECTION 17.2 MEASURING AND EXPRESSING ENTHALPY CHANGES

1. Calculate the enthalpy change for the reaction of 1.00 mol of H_2 with 1.00 mol of O_2 to form 2.00 mol of H_2O at 25.0°C. The enthalpy of formation of $\text{H}_2\text{O}(l)$ is -285.8 kJ/mol .

2. Calculate the enthalpy change for the reaction of 1.00 mol of H_2 with 1.00 mol of O_2 to form 2.00 mol of H_2O at 25.0°C. The enthalpy of formation of $\text{H}_2\text{O}(l)$ is -285.8 kJ/mol .

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7. Calculate the enthalpy change for the reaction of 1.00 mol of H_2 with 1.00 mol of O_2 to form 2.00 mol of H_2O at 25.0°C. The enthalpy of formation of $\text{H}_2\text{O}(l)$ is -285.8 kJ/mol .

8. Calculate the enthalpy change for the reaction of 1.00 mol of H_2 with 1.00 mol of O_2 to form 2.00 mol of H_2O at 25.0°C. The enthalpy of formation of $\text{H}_2\text{O}(l)$ is -285.8 kJ/mol .

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10. Calculate the enthalpy change for the reaction of 1.00 mol of H_2 with 1.00 mol of O_2 to form 2.00 mol of H_2O at 25.0°C. The enthalpy of formation of $\text{H}_2\text{O}(l)$ is -285.8 kJ/mol .

SECTION 17.3 HEAT IN CHANGES OF STATE

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Christian Drosten



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