

CHAPTER 15 REVIEW

Acid-Base Titration and pH**SECTION 1****SHORT ANSWER** Answer the following questions in the space provided.

1. Calculate the following values without using a calculator.

- _____ a. The $[\text{H}_3\text{O}^+]$ is 1×10^{-6} M in a solution. Calculate the $[\text{OH}^-]$.
- _____ b. The $[\text{H}_3\text{O}^+]$ is 1×10^{-9} M in a solution. Calculate the $[\text{OH}^-]$.
- _____ c. The $[\text{OH}^-]$ is 1×10^{-12} M in a solution. Calculate the $[\text{H}_3\text{O}^+]$.
- _____ d. The $[\text{OH}^-]$ in part c is reduced by half, to 0.5×10^{-12} M. Calculate the $[\text{H}_3\text{O}^+]$.
- _____ e. The $[\text{H}_3\text{O}^+]$ and $[\text{OH}^-]$ are _____ (directly, inversely, or not) proportional in any system involving water.

2. Calculate the following values without using a calculator.

- _____ a. The pH of a solution is 2.0. Calculate the pOH.
- _____ b. The pOH of a solution is 4.73. Calculate the pH.
- _____ c. The $[\text{H}_3\text{O}^+]$ in a solution is 1×10^{-3} M. Calculate the pH.
- _____ d. The pOH of a solution is 5.0. Calculate the $[\text{OH}^-]$.
- _____ e. The pH of a solution is 1.0. Calculate the $[\text{OH}^-]$.

3. Calculate the following values.

- _____ a. The $[\text{H}_3\text{O}^+]$ is 2.34×10^{-5} M in a solution. Calculate the pH.
- _____ b. The pOH of a solution is 3.5. Calculate the $[\text{OH}^-]$.
- _____ c. The $[\text{H}_3\text{O}^+]$ is 4.6×10^{-8} M in a solution. Calculate the $[\text{OH}^-]$.

PROBLEMS Write the answer on the line to the left. Show all your work in the space provided.4. $[\text{H}_3\text{O}^+]$ in an aqueous solution = 2.3×10^{-3} M.

- _____ a. Calculate $[\text{OH}^-]$ in this solution.

Chapter 15 Acid Base Titration Ph Section 2 Answers

Robert R. Redfield



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Table of Contents Chapter 15 Acid Base Titration Ph Section 2 Answers

1. Understanding the eBook Chapter 15 Acid Base Titration Ph Section 2 Answers
 - The Rise of Digital Reading Chapter 15 Acid Base Titration Ph Section 2 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 15 Acid Base Titration Ph Section 2 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 15 Acid Base Titration Ph Section 2 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 15 Acid Base Titration Ph Section 2 Answers
 - Personalized Recommendations
 - Chapter 15 Acid Base Titration Ph Section 2 Answers User Reviews and Ratings
 - Chapter 15 Acid Base Titration Ph Section 2 Answers and Bestseller Lists

5. Accessing Chapter 15 Acid Base Titration Ph Section 2 Answers Free and Paid eBooks
 - Chapter 15 Acid Base Titration Ph Section 2 Answers Public Domain eBooks
 - Chapter 15 Acid Base Titration Ph Section 2 Answers eBook Subscription Services
 - Chapter 15 Acid Base Titration Ph Section 2 Answers Budget-Friendly Options
6. Navigating Chapter 15 Acid Base Titration Ph Section 2 Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 15 Acid Base Titration Ph Section 2 Answers Compatibility with Devices
 - Chapter 15 Acid Base Titration Ph Section 2 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 15 Acid Base Titration Ph Section 2 Answers
 - Highlighting and Note-Taking Chapter 15 Acid Base Titration Ph Section 2 Answers
 - Interactive Elements Chapter 15 Acid Base Titration Ph Section 2 Answers
8. Staying Engaged with Chapter 15 Acid Base Titration Ph Section 2 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 15 Acid Base Titration Ph Section 2 Answers
9. Balancing eBooks and Physical Books Chapter 15 Acid Base Titration Ph Section 2 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 15 Acid Base Titration Ph Section 2 Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 15 Acid Base Titration Ph Section 2 Answers
 - Setting Reading Goals Chapter 15 Acid Base Titration Ph Section 2 Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 15 Acid Base Titration Ph Section 2 Answers
 - Fact-Checking eBook Content of Chapter 15 Acid Base Titration Ph Section 2 Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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