

Chapter 10 Chemical Quantities Review

Part 1: Vocabulary

Label the structure corresponding to the term in the vocabulary list.

- | | |
|-------------------------|---------------------------|
| _____ mole | _____ percent composition |
| _____ empirical formula | _____ empirical formula |
| _____ molar mass | _____ molar mass |
| _____ molar mass | _____ molar mass |

- _____ is a description of the relative amounts of each element in a compound.
- _____ is the lowest whole number ratio of the atoms of the elements in a compound.
- _____ is the mass of one mole of any gas at STP.
- _____ is the mass of one mole of a compound.
- _____ is the mass of one mole of a substance.
- _____ is the mass of one mole of a gas at STP.
- _____ is the mass of one mole of a gas at STP.

Part 2: Section Concept Review

10.1 The Mole: A Fundamental

Complete the following practice problems using your knowledge of the terms and your understanding of the concepts involved with each of the terms.

- _____ is the amount of substance in a sample, and it is related to a sample's quantity called the _____.
- _____ is the mass of one mole of a compound, substance, or element.
- _____ is the mass of one mole of a compound, substance, or element.

Practice Problems

- _____ is the mass of one mole of a substance.
- _____ is the mass of one mole of a substance.
- _____ is the mass of one mole of a substance.
- _____ is the mass of one mole of a substance.

Answers

Chapter 10 Chemical Quantities Guided Reading Answer Key

Hussin A.Rothana



Chapter 10 Chemical Quantities Guided Reading Answer Key:

Unveiling the Magic of Words: A Report on "**Chapter 10 Chemical Quantities Guided Reading Answer Key**"

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