



PENDIDIKAN KHAS (MASALAH PEMBELAJARAN) ENGLISH TAHUN 1

TOPIC : ADI'S FAMILY

Task : Drag the name of Adi's Family members to the correct picture.



father

mother

brother

sister

BY: CIKGU AMELIA

B006azqpr6 Bit1

R Pring



B006azqpr6 Bit1:

Unveiling the Power of Verbal Artistry: An Mental Sojourn through **B006azqpr6 Bit1**

In some sort of inundated with displays and the cacophony of immediate connection, the profound energy and psychological resonance of verbal beauty frequently fade into obscurity, eclipsed by the continuous assault of noise and distractions. Yet, nestled within the lyrical pages of **B006azqpr6 Bit1**, a captivating perform of fictional beauty that pulses with raw emotions, lies an remarkable journey waiting to be embarked upon. Composed by a virtuoso wordsmith, that enchanting opus manuals viewers on a psychological odyssey, lightly exposing the latent potential and profound influence stuck within the complex internet of language. Within the heart-wrenching expanse of this evocative examination, we can embark upon an introspective exploration of the book is central styles, dissect their interesting publishing type, and immerse ourselves in the indelible impression it leaves upon the depths of readers souls.

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8.3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation $8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} \times 18 \text{ g } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$ is your theoretical yield so percent yield $\frac{137.1 \text{ g } \text{H}_2\text{O}}{149.4 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$ problems 6 & 8 are similar to problem 5

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 160.0 g Al $\times \frac{1 \text{ mol Al}_2\text{O}_3}{2 \text{ mol Al}} \times 101.96 \text{ g } \text{Al}_2\text{O}_3 = 81.2 \text{ g } \text{Al}_2\text{O}_3$ c determine the percent yield $\frac{260.0 \text{ g}}{81.2 \text{ g}} \times 100 = 320.1\%$ theoretical 302.3 g

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web 7 Na₂C₂O₄ is the limiting reactant percent yield 86.68 only four molecules can be made 9 this amount cannot be

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