

## FORMULA - SHEET

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①  $PV = nRT$

P = Pressure (atm)  
V = Volume (Litre)  
n = mole  
R = gas constant (0.0821 atm L/mol K)  
T = Temperature (Kelvin 'K')

②  $d_{\text{gas}} = \frac{PM}{RT}$

M = Molar mass  
Baki salt upar hai

③  $\text{Density} = \frac{\text{mass}}{\text{volume}}$

④  $\frac{F - 32}{9} = \frac{C}{5} = \frac{K - 273}{5}$

F = Degree Fahrenheit  
C = Degree Centigrade  
K = Degree Kelvin

⑤  $1\text{m}^3 = 1000\text{L}$   
 $1\text{dm}^3 = 1\text{L}$   
 $1\text{cm}^3 = 1\text{ml}$

⑥  $1\text{amu} = \frac{1}{12} \times \text{mass of one } ^{12}\text{C isotope}$

$1\text{amu} = 1.66 \times 10^{-24}\text{gm}$

$1\text{amu} = \frac{1}{N_A}\text{gm}$

⑦  $\text{mole} = \frac{\text{mass (gm)}}{\text{molar mass}}$

$\text{mass} = \text{mole} \times \text{molar mass}$

⑧  $\text{mole} = \frac{\text{no. of particles (atom/molecule)}}{6.022 \times 10^{23}}$

$\text{no. of particles} = \text{mol} \times N_A$

⑨  $\text{mole} = \frac{\text{volume of gas at NTP/STP}}{22.4}$

$V_{\text{gas}} = \text{mol} \times 22.4$

⑩ मूल मंत्र

"Kuch bhi diya ho, sabse pehle mole nikalo fir jo puchhe, wo mole ki help se nikalo"

$\text{mole} = \frac{W}{M} = \frac{N}{N_A} = \frac{V_L}{22.4}$

# Assessment Of Basic Chemistry Concepts Answer Sheet

**Rachel Sandford**



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