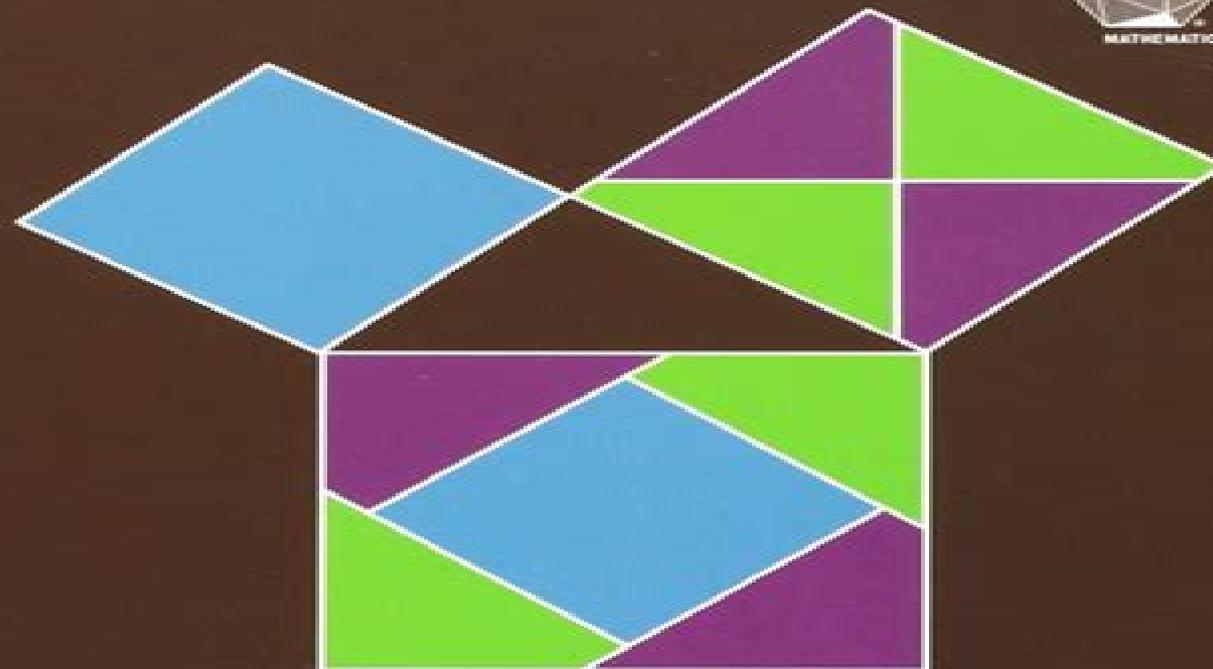




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web chemical equation in which all dissolved ionic reactants and products including spectator ions are explicitly represented by formulas for their dissociated ions molecular equation chemical equation in which all reactants and products are

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equation

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web mar 11 2014 solution ionic equation $\text{H}^+ \text{Cl}^- \text{Na}^+ \text{HCO}_3^- \text{Na}^+ \text{Cl}^- \text{H}_2\text{O} \text{CO}_2 \text{g}$ the spectator ions are Cl^- and Na^+
 net ionic equation $\text{H}^+ \text{Cl}^- \text{Na}^+ \text{HCO}_3^- \text{Na}^+ \text{Cl}^- \text{H}_2\text{O} \text{CO}_2 \text{g}$ $\text{H}^+ \text{HCO}_3^- \text{H}_2\text{O} \text{CO}_2 \text{g}$ here is a video on spectator ions
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web the six NO_3^- ions and the six Na^+ ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation $3\text{Ba}^{2+} + 2\text{PO}_4^{3-} \rightarrow \text{Ba}_3\text{PO}_4\text{2}$

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web these spectator ions whose presence is required to maintain charge neutrality are neither chemically nor physically changed by the process and so they may be eliminated from the equation to yield a more succinct representation called a

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 $\text{AgNO}_3 \text{KCl} \text{AgCl} \text{KNO}_3$ total ionic $\text{Ag}^+ \text{NO}_3^- \text{K}^+ \text{Cl}^- \text{AgCl} \text{K}^+ \text{NO}_3^-$ net ionic $\text{Ag}^+ \text{Cl}^- \text{AgCl}$

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water $\text{H}^+ \text{OH}^- \text{H}_2\text{O}$
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