

Chapter 17: Thermochemistry*Section 1: The Flow of Energy – Heat and Work*

1. What is thermochemistry? **The study of the energy changes that occur in changes of state or during chemical reactions.**
2. Heat always flows from a **warmer** object to a **cooler** object.
3. Given that $1 \text{ cal} = 4.184 \text{ J}$:
 - a. How many calories are in 325.9 J ? **77.89 cal**
 - b. How many J are in 34.7 cal ? **145.18 J**
4. What is the Law of Conservation of Energy (aka the First Law of Thermodynamics)? **States that in chemical and physical processes, energy is neither created nor destroyed – it is transferred or conserved.**
5. In an exothermic reaction:
 - a. Heat is **released** / absorbed
 - b. q is **positive** / **negative**
 - c. ΔH is **positive** / **negative**
 - d. Heat is a **reactant** / **product**
6. In an endothermic reaction:
 - a. Heat is **released** / **absorbed**
 - b. q is **positive** / **negative**
 - c. ΔH is **positive** / **negative**
 - d. Heat is a **reactant** / **product**
7. Diagram **A** represents an exothermic reaction, whereas diagram **B** represents an endothermic reaction:
8. What is the difference between heat capacity and specific heat capacity? **Specific heat is the amount of energy to increase the temperature of 1g of substance by 1°C**
9. The temperature of a 34.1 g sample of a substance increases from 29°C to 51°C when the sample absorbs 912 J of heat. What is the specific heat of the substance? **$1.126 \text{ J/g}^\circ\text{C}$**
10. When 867 J of heat are added to 9.3 g of a substance at STP, the temperature increases from 0°C to 14°C . What is the specific heat of the substance? **$6.66 \text{ J/g}^\circ\text{C}$**

Chapter 17 Thermochemistry Section Review Answers

Ralph H. Petrucci



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General Chemistry Key Concise Review Guide Sixth Edition Darrell D. Ebbing,1999 AP Chemistry Albert S. Tarendash,Frederick J. Rowe,2002 This indispensable guide to chemistry helps students who wish to prepare for the AP Chemistry exam on their own Comprehensive and easy to understand this learning guide includes a full content review two full length practice tests with hundreds of practice questions and thorough answer explanations and proven test taking strategies

Organic Chemistry T. W. Graham Solomons,Craig B. Fryhle,Scott A. Snyder,2016-01-19 The 12th edition of Organic Chemistry continues Solomons Fryhle Snyder s tradition of excellence in teaching and preparing students for success in the organic classroom and beyond A central theme of the authors approach to organic chemistry is to emphasize the relationship between structure and reactivity To accomplish this the content is organized in a way that combines the most useful features of a functional group approach with one largely based on reaction mechanisms The authors philosophy is to emphasize mechanisms and their common aspects as often as possible and at the same time use the unifying features of functional groups as the basis for most chapters The structural aspects of the authors approach show students what organic chemistry is Mechanistic aspects of their approach show students how it works And wherever an opportunity arises the authors show students what it does in living systems and the physical world around us

Complete Preparation for the MCAT Williams & Wilkins Review,1998-04 Here is the most respected test prep book for the Medical College Admission Test you can buy featuring an active learning approach for a better understanding of the exam s content and a better chance for success Unique to this guide are coverage of all recent changes in the MCAT plus a step by step plan for sharpening cognitive skills developing problem solving skills and critical thinking This thorough guide replaces expensive test preparation courses while giving students exactly what they need to get ready for the MCAT

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