

Name:

Date:

Chapter 17 & 19 Study Guide

Key Terms

Directions: Define each of the following key terms from chapters seventeen and nineteen on a separate sheet of paper.

- **Atom** - the smallest piece of matter that still has the properties of the element
- **atomic number** – number of protons in an atom's nucleus
- **atomic mass** – average mass of all an element's isotopes
- **ductile** – capable of being drawn into wire
- **electron** – subatomic particle with a negative charge
- **electron cloud** – area outside the nucleus in which electrons can be found.
- **Electron dot diagram** – uses the element's symbol and dots to represent the valence electrons in an element.
- **energy level** – areas within the electron cloud where electrons are most likely to be found.
- **Family** – vertical columns of the periodic table
- **Group** - vertical columns of the periodic table
- **half-life** – the amount of time it takes half the sample of a radioactive isotope to decay.
- **Isotope** – atoms of the same element that have different numbers of neutrons.
- **Lustrous** – having a shiny appearance
- **Malleable** – can be hammered flat or rolled into sheets
- **mass number** – sum of protons and neutrons in the nucleus of an atom.
- **Metal** – element that typically is hard, shiny, solid, malleable, and is a good conductor of heat and electricity.
- **Neutron** – neutrally charged subatomic particle.
- **Nonmetal** - element that typically is brittle solid or gas, not shiny, not malleable, and is a poor conductor of heat and electricity.
- **Nucleus** – center of the atom
- **Period** – horizontal row of the periodic table.
- **periodic table** – organized list of all known elements.
- **Proton** – positively charged subatomic particle
- **transition metal** – Elements found in families 3-12 of the periodic table.
- **transuranium element** – an element containing more than 92 protons in its nucleus
- **quark** – particle believed to make up protons and neutrons.
- **valence electron** – electrons found in the outermost energy level of an atom.

Chapter 17 Study Guide Answers

Stanley Chernicoff



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