

CHAPTER 14

FINANCIAL STATEMENT ANALYSIS

SUMMARY OF QUESTIONS BY LEARNING OBJECTIVES AND BLOOM'S TAXONOMY

Item	LO	BT	Item	LO	BT	Item	LO	BT	Item	LO	BT	Item	LO	BT
True-False Statements														
1.	1	C	9.	3	K	17.	4	K	25.	5	C	^{ns} 33.	4	K
2.	1	K	10.	3	AP	18.	5	K	26.	5	C	^{ns} 34.	5	K
3.	1	K	11.	3	C	19.	5	K	27.	6	K	^{ns} 35.	5	K
4.	1	K	12.	4	K	20.	5	K	28.	6	K	^{ns} 36.	6	K
5.	1	K	13.	4	C	21.	5	K	29.	7	K			
6.	2	K	14.	4	C	22.	5	C	30.	7	K			
7.	2	K	15.	4	C	23.	5	C	^{ns} 31.	2	K			
8.	3	K	16.	4	C	24.	5	C	^{ns} 32.	3	K			
Multiple Choice Questions														
37.	1	K	65.	4	C	93.	5	C	121.	5	AP	149.	6	AP
38.	1	K	66.	4	C	94.	5	C	122.	5	AP	150.	6	K
39.	1	C	67.	4	C	95.	5	C	123.	5	AP	151.	6	K
40.	1	C	68.	4	C	96.	5	AP	124.	5	AP	152.	6	AP
41.	1	C	69.	4	C	97.	5	C	125.	5	AP	153.	6	AP
42.	1	C	70.	5	C	98.	5	C	126.	5	AP	154.	6	C
43.	1	C	71.	5	C	99.	5	K	127.	5	AP	155.	6	K
44.	1	K	72.	5	K	100.	5	C	128.	5	AP	156.	6	K
45.	2	K	73.	5	K	101.	5	AP	129.	5	AP	157.	6	K
46.	2	K	74.	5	K	102.	5	AP	130.	5	AP	158.	6	C
47.	2	K	75.	5	AP	103.	5	C	131.	5	AP	159.	6	K
48.	2	K	76.	5	AP	104.	5	K	132.	5	AP	160.	7	C
49.	2	K	77.	5	AP	105.	5	AP	133.	5	AP	^{ns} 161.	1	K
50.	2	K	78.	5	AP	106.	5	C	134.	5	AP	^{ns} 162.	1	K
51.	3	K	79.	5	K	107.	5	C	135.	5	AP	^{ns} 163.	3	AN
52.	3	K	80.	5	C	108.	5	AP	136.	5	AP	^{ns} 164.	4	K
53.	3	K	81.	5	C	109.	5	K	137.	5	AP	^{ns} 165.	4	K
54.	3	AP	82.	5	K	110.	5	C	138.	5	AP	^{ns} 166.	5	K
55.	3	K	83.	5	AP	111.	5	K	139.	5	AP	^{ns} 167.	5	K
56.	3	C	84.	5	AP	112.	5	K	140.	5	AP	^{ns} 168.	5	K
57.	3	C	85.	5	AP	113.	5	K	141.	5	AP	^{ns} 169.	5	K
58.	3	C	86.	5	C	114.	5	AP	142.	5	AP	^{ns} 170.	6	K
59.	3	AP	87.	5	K	115.	5	AP	143.	5	AP	^{ns} 171.	6	AP
60.	3	AP	88.	5	AP	116.	5	AP	144.	5	AP	^{ns} 172.	6	K
61.	4	AP	89.	5	K	117.	5	AP	145.	6	AN			
62.	4	AP	90.	5	C	118.	5	AP	146.	6	C			
63.	4	K	91.	5	C	119.	5	AP	147.	6	AP			
64.	4	K	92.	5	C	120.	5	AP	148.	6	AP			

^{ns} This question also appears in the Study Guide.

ⁿ This question also appears in a self-test at the student companion website.

Chapter 14 Accounting Test Answers

LL Leslie



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web $m \times c \times \Delta t$ equation 1 where m is the total mass of the solution solute plus solvent c is the specific heat of the solution and Δt is the observed temperature change the specific heat of the solution is generally assumed to be the same as that of

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web the heat of reaction ΔH_{soln} is written after the products in units of kJ mol rxn for example sodium hydroxide dissolves exothermically releasing 44.2 kilojoules per mole dissolved this process is represented as $\text{NaOH (s)} \rightarrow \text{NaOH (aq)}$ $\Delta H_{\text{soln}} = -44.2 \text{ kJ mol}^{-1}$ write equations to similarly represent the dissolving process for

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web hand warmers are usually placed inside the gloves to keep your hands warm the heat keeping your hands warm is formed as a result of an exothermic reaction which occurs in a small packet hand warmers have 5 main ingredients in them salt carbon polypropylene vermiculite water and iron

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web when the hand warmer pack is activated the solid dissolves in water and produces a large temperature change in the form of heat in this laboratory experiment the heat of solution will be calculated from the data for six salts in order to

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web background the hand warmers are filled with sodium acetate click the metal disc and you will notice the liquid changes state and becomes a solid in the process heat is released the sodium acetate is a supersaturated solution

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web the goal of my experiment and research is to evaluate the various number of compounds to be used as hand warmers and locate the ones that can raise the temperature to 20 degrees celsius in 100 g water and costs as little as possible

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web in this lab you are challenged to use chemistry to determine which ionic compound makes the best hand warmer in this experiment you will collect data that will allow you to calculate the heat of solution ΔH_{soln} or the amount of energy released or absorbed in kJ mol of the substance being dissolved

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web mar 4 2014 contains water the hand warmer is struck in a manner that ruptures the inner pouch releasing the ionic salt into the water of the outer pouch the salt dissolves and the water warms this type of hand warmer tends to produce a more vigorous heat than the dry powder type of hand warmer but does not produce heat for quite as long

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web lab investigation in today s experiment you will design and execute an experimental procedure to determine which of three ionic compounds is most suitable for use in a hand warmer the three compounds you will use is determined by the group number you are assigned the ideal hand warmer

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