

CHAPTER 13

Nuclei

In the last chapter, we learnt that atomic nucleus was discovered in the year 1911 by Rutherford and his associates from the study of large angle scattering of alpha particles from the thin metal foil. In this chapter, we shall study the constituents of the nucleus and how they are held together. We shall also study different properties of nuclei such as size, mass, density and stability of nuclei and associated phenomena such as radioactivity, nuclear fission and nuclear fusion.

NUCLEUS

As per the Rutherford's experiment, it was concluded that the entire positive charge and mass of an atom was present at its very tiny centre known as nucleus.

Composition of nucleus

Nucleus is composed of two types of particles

- (i) **Protons** Positively charged particles
- (ii) **Neutrons** Neutral particles

Mass of neutron is slightly greater than the mass of the proton.

$$\text{Mass of proton, } m_p = 1.6726231 \times 10^{-27} \text{ kg}$$

$$\text{Mass of neutron, } m_n = 1.6749286 \times 10^{-27} \text{ kg}$$

$$m_p \approx 1.8-2.0 m_e, \text{ where } m_e = \text{mass of electron.}$$

Atomic number

Atomic number of an element is the number of proton present inside the nucleus of a neutral atom. It is also equal to the number of electrons revolving in various orbits around the nucleus of an uncharged atom. It is represented by Z .

Atomic number (Z) = Number of proton = Number of electron

Mass number

The total number of neutrons and protons, i.e. number of nucleons present inside the nucleus of an element is called its mass number A . Number of neutrons is represented by N .

So, **Mass number (A) = $Z + N$**

A chemical element is represented by ${}_Z$ (chemical symbol) ^{A} or ${}_Z X^A$.

Total charge of the nucleus is equal to the total charge on all protons Z in it, i.e. Ze , where $e = 1.6 \times 10^{-19} \text{ C}$.

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