

## AP Chemistry Chapter 4 Practice Exam

D

1. The Claus reactions, shown below, are used to generate elemental sulfur from hydrogen sulfide.

How much sulfur (in grams) is produced from 21.5 grams of  $\text{O}_2$ ?

A) 97.0 g

B) 14.4 g

C) 21.5 g

D) 43.1 g

E) none of these

$$\frac{21.5 \text{ g O}_2}{32.00 \text{ g O}_2} \times \frac{1 \text{ mol O}_2}{3 \text{ O}_2} \times \frac{2 \text{ SO}_2}{1 \text{ SO}_2} \times \frac{3 \text{ S}}{1 \text{ SO}_2} \times \frac{32.06 \text{ g S}}{1 \text{ mol S}} = 43.1 \text{ g S}$$

B

2. Iron is produced from its ore by the reactions:

How many moles of  $\text{O}_2(g)$  are needed to produce 4.6 moles of  $\text{Fe}(s)$ ?A) 2.3 mole  $\text{O}_2$ B) 3.5 mole  $\text{O}_2$ C) 4.6 mole  $\text{O}_2$ D) 6.9 mole  $\text{O}_2$ E) 13.8 mole  $\text{O}_2$ 

$$\frac{4.6 \text{ mol Fe}(s)}{2 \text{ Fe}} \times \frac{3 \text{ CO}}{2 \text{ CO}} \times \frac{1 \text{ O}_2}{2 \text{ CO}} = 3.45 \text{ mol O}_2$$

B

3. Suppose the reaction
- $\text{Ca}_3(\text{PO}_4)_2 + 3\text{H}_2\text{SO}_4 \rightarrow 3\text{CaSO}_4 + 2\text{H}_3\text{PO}_4$
- is carried out starting with 153 g of
- $\text{Ca}_3(\text{PO}_4)_2$

and 76.8 g of  $\text{H}_2\text{SO}_4$ . How much phosphoric acid will be produced?

A) 76.7 g

B) 31.1 g

C) 229.8 g

D) 115.1 g

E) 96.7 g

$$\frac{153 \text{ g Ca}_3(\text{PO}_4)_2}{326 \text{ g Ca}_3(\text{PO}_4)_2} \times \frac{1 \text{ mol Ca}_3(\text{PO}_4)_2}{1 \text{ Ca}_3(\text{PO}_4)_2} \times \frac{2 \text{ H}_3\text{PO}_4}{1 \text{ Ca}_3(\text{PO}_4)_2} =$$

$$\frac{76.8 \text{ g H}_2\text{SO}_4}{98 \text{ g H}_2\text{SO}_4} \times \frac{1 \text{ mol H}_2\text{SO}_4}{3 \text{ H}_2\text{SO}_4} \times \frac{2 \text{ H}_3\text{PO}_4}{1 \text{ H}_2\text{SO}_4} =$$

B

- 4.
- $\text{SO}_2$
- reacts with
- $\text{H}_2\text{S}$
- as follows:

When 7.50 g of  $\text{H}_2\text{S}$  reacts with 12.75 g of  $\text{SO}_2$ , which statement applies?

A) 6.38 g of sulfur are formed.

B) 10.6 g of sulfur are formed.

C) 0.0216 moles of  $\text{H}_2\text{S}$  remain.D) 1.13 g of  $\text{H}_2\text{S}$  remain.E)  $\text{SO}_2$  is the limiting reagent.

$$\frac{7.50 \text{ g H}_2\text{S}}{34.03 \text{ g H}_2\text{S}} \times \frac{1 \text{ mol H}_2\text{S}}{2 \text{ H}_2\text{S}} \times \frac{3 \text{ S}}{1 \text{ H}_2\text{S}} \times \frac{32 \text{ g S}}{1 \text{ mol S}} = 10.89$$

$$\frac{12.75 \text{ g SO}_2}{64 \text{ g SO}_2} \times \frac{1 \text{ mol SO}_2}{1 \text{ SO}_2} \times \frac{3 \text{ S}}{1 \text{ SO}_2} \times \frac{32 \text{ g S}}{1 \text{ mol S}} = 19.13$$

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## **Table of Contents Ap Chem Chapter 1 Practice Test**

1. Understanding the eBook Ap Chem Chapter 1 Practice Test
  - The Rise of Digital Reading Ap Chem Chapter 1 Practice Test
  - Advantages of eBooks Over Traditional Books
2. Identifying Ap Chem Chapter 1 Practice Test
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ap Chem Chapter 1 Practice Test
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ap Chem Chapter 1 Practice Test
  - Personalized Recommendations
  - Ap Chem Chapter 1 Practice Test User Reviews and Ratings
  - Ap Chem Chapter 1 Practice Test and Bestseller Lists
5. Accessing Ap Chem Chapter 1 Practice Test Free and Paid eBooks

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  - Ap Chem Chapter 1 Practice Test eBook Subscription Services
  - Ap Chem Chapter 1 Practice Test Budget-Friendly Options
6. Navigating Ap Chem Chapter 1 Practice Test eBook Formats
- ePub, PDF, MOBI, and More
  - Ap Chem Chapter 1 Practice Test Compatibility with Devices
  - Ap Chem Chapter 1 Practice Test Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Ap Chem Chapter 1 Practice Test
  - Highlighting and Note-Taking Ap Chem Chapter 1 Practice Test
  - Interactive Elements Ap Chem Chapter 1 Practice Test
8. Staying Engaged with Ap Chem Chapter 1 Practice Test
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ap Chem Chapter 1 Practice Test
9. Balancing eBooks and Physical Books Ap Chem Chapter 1 Practice Test
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ap Chem Chapter 1 Practice Test
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Ap Chem Chapter 1 Practice Test
- Setting Reading Goals Ap Chem Chapter 1 Practice Test
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ap Chem Chapter 1 Practice Test
- Fact-Checking eBook Content of Ap Chem Chapter 1 Practice Test
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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