

Soluzioni

CAPITOLO 15

VERIFICA LE TUE CONOSCENZE

VALENZA E NUMERO DI OSSIDAZIONE

- 1**
- La valenza di un elemento è il numero di elettroni che l'atomo cede, acquista o condivide quando si lega ad altri atomi.
 - La valenza massima corrisponde al numero romano che contraddistingue il gruppo, quindi, nel caso di un elemento del terzo gruppo della tavola periodica, è 3.
 - Un atomo di carbonio può formare al massimo 4 legami.
- 2** (C) di cariche reali o convenzionali associato a ciascun atomo del composto.
- 3** Regole per il calcolo del numero di ossidazione non corrette [tra parentesi, la versione corretta]
- l'ossigeno ha sempre n.o. -2 ; [l'ossigeno nei composti ha n.o. -2 , tranne nei perossidi, nei superossidi e nel difluoruro di ossigeno, dove ha rispettivamente n.o. -1 , $-1/2$, $+2$];
 - l'idrogeno ha n.o. $+1$ nei composti e $+2$ nella molecola H_2 ; [l'idrogeno nei composti ha n.o. $+1$ quando è legato a un non metallo e -1 quando è legato a un metallo].

SCRIVERE LE FORMULE DEI COMPOSTI A PARTIRE DAI NUMERI DI OSSIDAZIONE

- 4** Na_2O : in questo modo la somma algebrica dei n.o. è uguale a zero: $2 \cdot (+1) + (-2) = 0$.
- 5** $Ca_3(PO_4)_2$
- 6**
- SO_4^{2-} ; carica dello ione = -2
 - ClO_7^- ; carica dello ione = -1
- 7** La formula Cl_2Ba è sbagliata, in quanto si scrive a sinistra l'atomo con n.o. positivo (Ba_2Cl_2).
- 8** È corretta soltanto: d. SO_2 [tra parentesi, le formule corrette]
- HS (H_2S)
 - Fe_2S_2 (Fe_2S_3)
 - LiS_2 (Li_2S)

LA NOMENCLATURA DEI COMPOSTI BINARI

- 9** Un composto binario è formato da due elementi diversi. I composti binari si dividono in:
- composti contenenti ossigeno: ossidi (es. ZnO), perossidi (es. H_2O_2), superossidi;
 - composti contenenti idrogeno: idruri (es. NaH), idracidi (es. HF);
 - sali binari (es. $NaCl$, $CaCl_2$).
- 10** In molti composti binari si ha la desinenza *-ure*; tale desinenza non compare negli ossidi.
- 11** La notazione di Stock si utilizza quando gli elementi che formano un composto hanno più di un numero di ossidazione. Comporta l'eliminazione dei prefissi numerici e l'introduzione del numero di ossidazione dell'elemento, come numero romano tra parentesi posto dopo il nome dell'elemento. Si ritiene che in questo modo il nome che contraddistingue l'elemento risulti più breve e più chiaro.
- | | | |
|-------------------|----------|-----------------------|
| (n.o. Fe = $+2$) | $FeBr_2$ | bromuro di ferro(II) |
| (n.o. Fe = $+3$) | $FeBr_3$ | bromuro di ferro(III) |
- 12** Lo ione ossido (O^{2-}) è formato da un atomo di O con n.o. = -2 .
Lo ione perossido (O_2^{2-}) è formato da due atomi di O entrambi con n.o. = -1 , legati da legame covalente.
Lo ione superossido (O_2^-) è formato da due atomi di O con n.o. medio = $-1/2$, legati da legame covalente.
Un ossido acido è formato da un non metallo e ossigeno, un ossido basico è formato da un metallo e ossigeno.
- 13**
- Idrogeno e metallo: idruri (NaH , LiH);
 - idrogeno e non metallo dei gruppi 16 e 17: idracidi (HCl , HF , HBr);
 - metallo e non metallo: sali binari ($NaCl$, $FeCl_2$).
- 14** (B) cianidrico.
- 15** K_2O : ossido di potassio;
 SO_2 : anidride solforosa (n.o. S = $+4$).
 K_2O si chiama *ossido* perché è un ossido basico (metallo e ossigeno).
 SO_2 si chiama *anidride* perché è un ossido acido (non metallo e ossigeno).

Soluzioni Libro Zanichelli Chimica

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Non avere paura di sognare Alberto Mantovani,2016-09-01T00:00:00+02:00 Quando l'astronauta Neil Armstrong ha lasciato la prima impronta dell'uomo sulla Luna la sua voce entrata in tutte le case Un piccolo passo per un uomo un grande salto per l'umanità La scienza si racconta attraverso le intuizioni delle menti geniali che la illuminano ma spesso lascia nell'ombra i piccoli passi necessari a raggiungere risultati importanti Alberto Mantovani il ricercatore italiano piú citato al mondo ma soprattutto uno scienziato che ama il suo lavoro Leggendo questo decalogo per aspiranti scienziati ci si imbatte in scoperte impreviste incontri emozionanti e pagina dopo pagina si scopre una guida preziosa per realizzare se stessi anche al di fuori del mondo della ricerca Le storie di questo libro parlano di passione e rispetto creativit e apertura al mondo invitano a non temere la fatica ad ascoltare gli altri e a dare valore alle loro critiche Un decalogo per affrontare le sfide nel lavoro e nella vita con l'ambizione e il coraggio di cambiare in meglio il mondo in cui viviamo *L'Italia che scrive rassegna per coloro che leggono* ,1922 L'Italia che scrive ,1921 Rivista scientifico-industriale delle principali scoperte ed invenzioni fatte nelle scienze e nelle industrie ,1908 Cosa significa davvero studiare chimica all'università? Valentina Domenici,Antonella Maggio,Laura Orian,2025-12-09T00:00:00+01:00 Una raccolta di testimonianze e riflessioni dedicate allo studio della Chimica all'università Il volume aiuta studenti e studentesse a orientarsi nella scelta del percorso offrendo uno sguardo diretto sull'esperienza formativa sul metodo di studio e sulle motivazioni che guidano la decisione *Storia del libro napoletano. Volume primo. Epoche - Cultura - Testi* Lucio Pironti,2023-12-22 Monumenti strade castelli musei sono gli immediati indicatori per chiunque voglia farsi un'idea della civiltà di Napoli ma nelle biblioteche e piú precisamente nel libro sia manoscritto che stampato che quelle indicazioni trovano una spiegazione piú compiuta nelle pagine degli autori che hanno scritto sugli innumerevoli aspetti della città Tale concetto costituisce lo sfondo di questa Storia del libro napoletano cercando di far emergere le peculiarità piú rilevanti della storia e della cultura napoletana attraverso le testimonianze librarie piú importanti che partono dai papiri ercolanesi per seguire con i codici medievali manoscritti talora splendidi per le illuminazioni fino ad arrivare al 1471 quando si ebbe la prima stampa a caratteri mobili il Bartolo di Sassoferrato *Lectura* un testo giuridico che inaugura una caratteristica non secondaria della cultura napoletana per concludere al 1860 La

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