

# Successful PROJECT MANAGEMENT

SEVENTH EDITION



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# Project Management 7th Edition Solution Manual

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web explain how a molecule that contains polar bonds can be nonpolar answer problem 6 2 2 6 2 2 which of the following molecules and ions contain polar bonds which of these molecules and ions have dipole moments a  $\text{ClF}_5$  b  $\text{ClO}_2$  c  $\text{TeCl}_4$  d  $\text{PCl}_3$  e  $\text{SeF}_4$  f  $\text{PH}_2$  g  $\text{XeF}_2$  answer problem 6 2 3 6 2 3

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web the relative electronegativities of the bonded atoms is  $\text{H} < \text{C} < \text{Cl}$  and so the bond moments all point toward the  $\text{Cl}$  end of the molecule and sum to yield a considerable dipole moment the molecules are relatively polar for molecules of high symmetry such as  $\text{BF}_3$  trigonal planar  $\text{CH}_4$  tetrahedral  $\text{PF}_5$  trigonal bipyramidal and  $\text{SF}_6$

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web explain the difference between a nonpolar covalent bond a polar covalent bond and an ionic bond answer nonpolar covalent electronegativity difference is less than 0.4 nonmetal nonmetal close together on the periodic table

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web apr 1 2021 updated on april 01 2021 chemical bonds may be classified as being either polar or nonpolar the difference is how the electrons in the bond are arranged key takeaways what is a polar bond in chemistry a polar bond is a type of covalent bond in which the electrons forming the bond are unequally distributed