

Lab 22A

Investigating Chemical Equilibrium

Purpose: To investigate the macroscopic properties of two chemical systems at equilibrium, observe shifts in equilibrium concentrations as a function of concentration and temperature and relate the to systems, which are exploring the structures through applying Le Chatelier's principle.

Proble:

1. Write a balanced equation for the reaction in part 1.
 $\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$
2. Write a balanced equation for the reaction in part 2.
 $\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$
3. Write a balanced equation for the reaction in part 3.
 $\text{Co}(\text{H}_2\text{O})_6^{2+}(\text{aq}) + 4\text{Cl}^{-}(\text{aq}) \rightleftharpoons \text{CoCl}_4^{2-}(\text{aq}) + 6\text{H}_2\text{O}(\text{l})$
4. In part 4, what is the conjugate acid of the base $\text{Co}(\text{H}_2\text{O})_6^{2+}$ given as $\text{Co}(\text{H}_2\text{O})_6^{2+}$ in the lab? What is the conjugate base of the acid $\text{Co}(\text{H}_2\text{O})_6^{2+}$?
Conjugate acid: $\text{Co}(\text{H}_2\text{O})_7^{3+}$ (the central group and have the other bond for adding. The conjugate base: $\text{Co}(\text{H}_2\text{O})_5^{+}$ (the central is the proton and other bond must have the other bond).
5. What is the conjugate acid of the base $\text{Co}(\text{H}_2\text{O})_6^{2+}$?
Conjugate acid: $\text{Co}(\text{H}_2\text{O})_7^{3+}$ (the central is the proton and other bond must have the other bond).
6. What is the conjugate base of the acid $\text{Co}(\text{H}_2\text{O})_6^{2+}$?
Conjugate base: $\text{Co}(\text{H}_2\text{O})_5^{+}$ (the central is the proton and other bond must have the other bond).
7. Write the equation for the reaction between Fe^{3+} and OH^{-} ions.
 $\text{Fe}^{3+} + 3\text{OH}^{-} \rightleftharpoons \text{Fe}(\text{OH})_3$
8. Write the equation for the reaction between Fe^{3+} and OH^{-} ions.
 $\text{Fe}^{3+} + 3\text{OH}^{-} \rightleftharpoons \text{Fe}(\text{OH})_3$
9. Write the equation for the reaction between Fe^{3+} and OH^{-} ions.
 $\text{Fe}^{3+} + 3\text{OH}^{-} \rightleftharpoons \text{Fe}(\text{OH})_3$
10. Write the equation for the reaction between Fe^{3+} and OH^{-} ions.
 $\text{Fe}^{3+} + 3\text{OH}^{-} \rightleftharpoons \text{Fe}(\text{OH})_3$

Investigating Chemical Equilibrium Lab 12a Answer

M Walker



Investigating Chemical Equilibrium Lab 12a Answer:

Investigating Chemical Equilibrium Lab 12a Answer Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Investigating Chemical Equilibrium Lab 12a Answer**," published by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we will delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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