

## Unit One

## Lesson 1

1. 1. d    2. d    3. a    4. d    5. c    6. b  
 7. b    8. b    9. c    10. a    11. d    12. b  
 13. a    14. c    15. a    16. d    17. d    18. d  
 19. d    20. c    21. a    22. b    23. a    24. d  
 25. a    26. b    27. c    28. c    29. b    30. b

2. 1. c    2. d    3. a    4. b

3. 1. c.B    2. d.A    3. a.E    4. b.D

4. 1. (✓)  
 2. (✗) Mendeleev's periodic table is .....  
 3. (✗) ..... put in vertical groups.  
 4. (✗) ..... 67 elements.  
 5. (✗) ..... new period.  
 6. (✗) ..... ascendingly .....  
 7. (✓)    8. (✗) Mendeleev .....  
 9. (✗) ..... is 118 elements.  
 10. (✓)    11. (✗) ..... into 4 blocks.  
 12. (✓)    13. (✗) ..... on the right side .....  
 14. (✓)    15. (✗) ..... 6 groups.  
 16. (✗) ..... by (B).  
 17. (✗) Elements of groups (B) .....  
 18. (✗) ..... its period number.  
 19. (✗) The group number of .....  
 20. (✓)    21. (✗) ..... is 12  
 22. (✓)    23. (✗) ..... same period .....  
 24. (✓)

5. 1. Mendeleev's periodic table.  
 2. Moseley's periodic table.    3. Protons.  
 4. Modern periodic table.  
 5. Groups.    6. Periods.  
 7. Main energy levels.  
 8. Transition elements.    9. p-block.  
 10. Inert (noble) gases.    11. f-block.  
 12. Period number.    13. Group number.  
 14. Atomic number.

10. sublevels.  
 11. their atomic numbers – way of filling their energy sublevels with electrons.  
 12. 118 – 92 – artificially.    13. 7 – 18  
 14. left – 1A – 2A    15. right – six  
 16. 16 – 6    17. B – 8 – three  
 18. transition    19. 4 – 10  
 20. 18 – 15    21. left – right  
 22. period – group    23. 2 – 5A  
 24. 11    25. 3 – 3A  
 26. second – 1A    27. zero – 2  
 28. 12    29. group – period

## 7 ①

Element	Its atomic number	Its electronic configuration			Its position in periodic table	
		K	L	M	Period No.	Group No.
${}_6\text{C}$	6	2	4	–	2	(4A)
${}_8\text{O}$	8	2	6	–	2	(6A)
${}_{12}\text{Mg}$	12	2	8	2	3	(2A)
${}_{17}\text{Cl}$	17	2	8	7	3	(7A)

## ②

Element	${}_{20}\text{Ca}$	${}_{15}\text{P}$	${}_{10}\text{Ne}$	${}_1\text{H}$
Electronic configuration :				
No. of energy levels :	4	3	2	1
Period number :	4	3	2	1
No. of electrons in the outermost level :	2	5	8	1
Group number :	2A	5A	0	1A

8. 1. Mendeleev.    2. Rutherford.  
 3. & 4. Moseley.    5. Bohr.

9. 1. To facilitate their study and find a relationship between elements and their physical and chemical properties.  
 2. Because he predicted the discovery of new elements and determined the values of their

# El Moasser Answer

**O García**



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