

Specific Heat and Heat Capacity

Name _____ KEY _____ Period _____

1. Calculate the amount of heat needed to increase the temperature of 250 g of water from 20 °C to 46 °C.

$$q = m \times C \times \Delta T$$
$$q = 250 \text{ g} \times 4.18 \text{ J/g}^\circ\text{C} \times 26^\circ\text{C}$$
$$q = 27,170 \text{ J or } 27 \text{ kJ}$$

2. Calculate the specific heat capacity of copper given that 204.75 J of energy raises the temperature of 15 g of copper from 25 °C to 60. °C.

$$q = m \times C \times \Delta T$$
$$C = q/m \times \Delta T$$
$$C = 204.75 \text{ J} / (15 \text{ g} \times 35^\circ\text{C})$$
$$C = 0.39 \text{ J/g}^\circ\text{C}$$

3. 216 J of energy is required to raise the temperature of aluminum from 15 °C to 35 °C. Calculate the mass of aluminum. (Specific heat capacity of aluminum is 0.90 J/g °C).

$$q = m \times C \times \Delta T$$
$$m = q/C \times \Delta T$$
$$m = 216 \text{ J} / (0.90 \text{ J/g}^\circ\text{C} \times 20^\circ\text{C})$$
$$m = 12 \text{ g}$$

4. The initial temperature of 150 g of ethanol was 22 °C. What will be the final temperature of the ethanol if 3240 J was needed to raise the temperature of the ethanol? (Specific heat capacity of ethanol is 2.44 J/g °C).

$$q = m \times C \times \Delta T$$
$$\Delta T = q/m \times C$$
$$\Delta T = 3240 \text{ J} / (150 \text{ g} \times 2.44 \text{ J/g}^\circ\text{C})$$
$$\Delta T = 8.85^\circ\text{C}$$
$$T_{\text{final}} = 22^\circ\text{C} + 8.85^\circ\text{C} = 31^\circ\text{C}$$

5. When 435 J of heat is added to 3.4 g of olive oil at 21 °C, the temperature increases to 85 °C. What is the specific heat of the olive oil?

$$q = m \times C \times \Delta T$$
$$C = q/m \times \Delta T$$
$$C = 435 \text{ J} / (3.4 \text{ g} \times 64^\circ\text{C})$$
$$C = 2.0 \text{ J/g}^\circ\text{C}$$

6. How many kilojoules of heat are absorbed when 1.00 L of water is heated from 18 °C to 85 °C? (Hint: You first need to determine the mass of the water, then calculate q in the requested unit.)

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