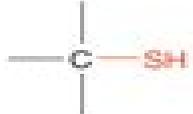
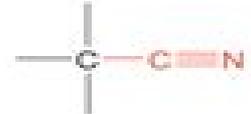
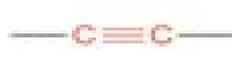
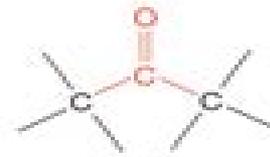
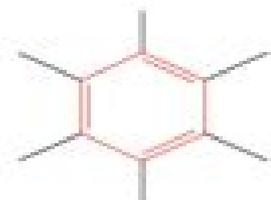
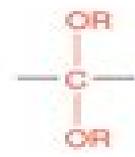
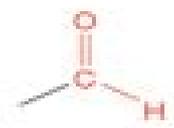
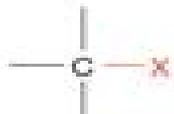
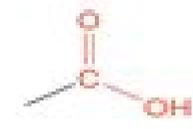
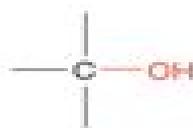
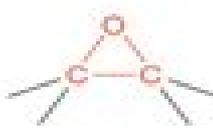
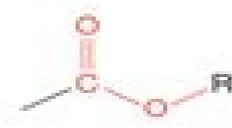
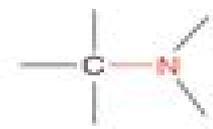
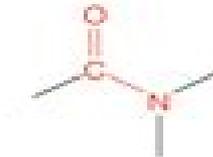


**TABLE 1-6 Common Functional Groups<sup>a</sup>**

Functional Group (Red)	Compound Class	Functional Group (Red)	Compound Class	Functional Group (Red)	Compound Class
	Alkene		Thiol		Nitrile
	Alkyne		Ether		Ketone
	Arene or Aromatic compound		Acetal		Aldehyde
 (X = F, Cl, Br, I)	Alkyl halide		Hemiacetal		Carboxylic acid
	Alcohol		Epoxide		Ester
	Phenol		Amine		Amide

<sup>a</sup>R indicates an alkyl group, and the absence of an atom at the end of a bond indicates that either R or H may be attached.

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web dec 26 2015 in the nuclear industry workers use a rule of thumb that the radioactivity from any sample will be relatively harmless after 10 half lives calculate the fraction of a radioactive sample that remains after this time period hint radioactive decay obeys first

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web in order to have successful collisions the colliding particles must have both the proper amount of energy and the proper fill in blank draw an energy diagram which shows the shape of the curve you might expect for the reaction in this question the overall reaction is exothermic

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web mar 13 2023 work in groups on these problems you should try to answer the questions without referring to your textbook if you get stuck try asking another group for help 1 for the reaction below if substance a is disappearing at a rate of  $1.82 \times 10^{-2} \text{ mol l}^{-1} \text{ s}^{-1}$  at what rate is c appearing 3 a 3 b  $\rightarrow$  5 c 2 d

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web kinetics practice problems and solutions d write the rate law for the overall reaction rate  $k_a = 2.9$  consider the following mechanism  $\text{o}_3 \rightarrow \text{o}_2 + \text{o}$  fast  $\text{o}_3 + \text{o}_2 \rightarrow \text{o}_2 + \text{o}_2$  slow a write the overall balanced chemical equation  $2 \text{o}_3 \rightarrow 3 \text{o}_2$  b identify any intermediates within the mechanism c what is the order with respect

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web a<sub>2</sub> 1 st b<sub>2</sub> 1 st kinetics practice problems and solutions d write the rate law for the overall reaction rate  $k_a = 2.9$  consider the following mechanism  $\text{o}_3 \rightarrow \text{o}_2 + \text{o}$  fast  $\text{o}_3 + \text{o}_2 \rightarrow \text{o}_2 + \text{o}_2$  slow a write the overall balanced chemical equation  $2 \text{o}_3 \rightarrow 3 \text{o}_2$  b identify any intermediates within the mechanism c

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web reaction rates google classroom mg s + 2 hcl a q  $\rightarrow$  mgcl<sub>2</sub> a 2 a q + h<sub>2</sub> g the reaction between mg s and hcl a q is represented by the equation above in a kinetics experiment a 0.080 g sample of mg s ribbon is added to an erlenmeyer flask containing 50 ml of 3.0 m hcl a q at 25 c

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