

Electrochemistry

Q1 Calculate the Emf of the following cell at 298K
 $\text{Fe(s)} | \text{Fe}^{2+} (0.01\text{M}) || \text{H}^+ (1\text{M}) | \text{H}_2(\text{g}) (1\text{bar}), \text{Pt}$
 $E^\circ_{\text{cell}} = 0.44\text{V}$

2023
(2 Marks)



Nernst equation

$$E_{\text{cell}} = E^\circ_{\text{cell}} - \frac{0.059}{n} \log \frac{[\text{Fe}^{2+}][\text{H}_2]}{[\text{H}^+]^2}$$
$$E_{\text{cell}} = 0.44 - \frac{0.059}{2} \log \frac{[10^{-2}][1]}{[1]^2}$$
$$E_{\text{cell}} = 0.44 - \frac{0.059}{2} \log 10^{-2}$$
$$E_{\text{cell}} = 0.44 + \frac{0.059}{2} \times 2$$
$$E_{\text{cell}} = \underline{0.499\text{V}}$$

Electrochemistry Problems And Solutions

Ying Liu



Electrochemistry Problems And Solutions:

Unveiling the Energy of Verbal Artistry: An Emotional Sojourn through **Electrochemistry Problems And Solutions**

In a global inundated with displays and the cacophony of fast interaction, the profound energy and mental resonance of verbal artistry frequently disappear into obscurity, eclipsed by the constant assault of noise and distractions. Yet, nestled within the musical pages of **Electrochemistry Problems And Solutions**, a captivating perform of fictional splendor that pulses with fresh emotions, lies an unforgettable trip waiting to be embarked upon. Published by a virtuoso wordsmith, this magical opus books viewers on an emotional odyssey, lightly exposing the latent potential and profound influence embedded within the delicate web of language. Within the heart-wrenching expanse with this evocative examination, we shall embark upon an introspective exploration of the book is central subjects, dissect their captivating writing design, and immerse ourselves in the indelible impact it leaves upon the depths of readers souls.

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