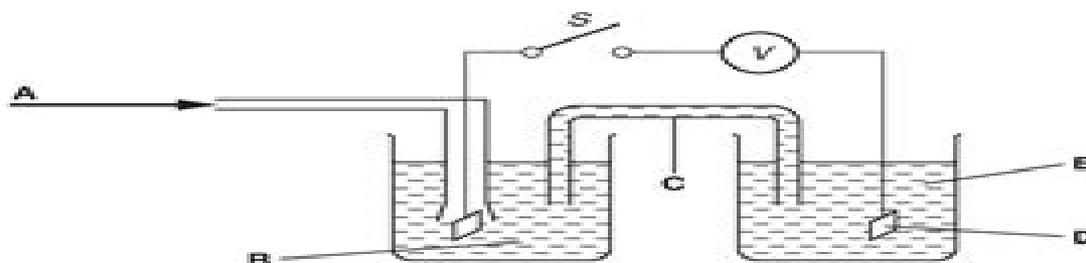


ELECTROCHEMISTRY WS 1

- 1 Chlorine gas and iron(II) ions react together in aqueous solution as follows.



- (a) The following diagram shows the apparatus needed to measure the E_{cell}° for the above reaction.



- (i) In the spaces below, identify what the five letters A – E in the above diagram represent.

A chlorine at $P_{\text{Cl}_2} = 1 \text{ atm}$ and 298 K
 B $1 \text{ mol dm}^{-3} \text{ NaCl}$
 C salt bridge
 D platinum
 E $1 \text{ mol dm}^{-3} \text{ Fe}^{2+} \text{ aq}$ and $1 \text{ mol dm}^{-3} \text{ Fe}^{3+} \text{ aq}$

- (ii) Use the *Data Booklet* to calculate the E_{cell}° for this reaction, and hence decide which direction (left to right, or right to left) electrons would flow through the voltmeter V when switch S is closed.



[7]

Electrochemistry Answers

Hugh Stott Taylor



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Solutions Manual to Accompany Elements of Physical Chemistry David Smith,2013-05-30 The Solutions Manual to accompany Elements of Physical Chemistry 6th edition contains full worked solutions to all end of chapter discussion questions and exercises featured in the book The manual provides helpful comments and friendly advice to aid understanding It is also a valuable resource for any lecturer who wishes to use the extensive selection of exercises featured in the text to support either formative or summative assessment and wants labour saving ready access to the full solutions to these questions

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