

1. How many electrons are needed to balance the half-reaction?



- A) 0 B) 6 C) 8 D) 4 E) 10

2. In an electrolytic cell, the positive electrode is the

- A) anode, where oxidation occurs
B) anode, where reduction occurs
C) cathode, where oxidation occurs
D) cathode, where reduction occurs

3. For which chemical reaction must an electrolytic cell be used?

- A) $\text{AgNO}_3 + \text{NaCl} \rightarrow \text{AgCl} + \text{NaNO}_3$
B) $\text{Cu} + \text{FeCl}_2 \rightarrow \text{CuCl}_2 + \text{Fe}$
C) $\text{Zn} + 2 \text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2$
D) $2 \text{Al} + 3 \text{Ni}(\text{NO}_3)_2 \rightarrow 2 \text{Al}(\text{NO}_3)_3 + 3 \text{Ni}$

4. Which occurs in an electrolytic cell containing CuCl_2 (aq)?

- A) Cu^{2+} ions migrate toward the positive electrode.
B) Cl^- ions migrate toward the negative electrode.
C) Cu^{2+} ions are reduced.
D) Cl^- ions are reduced.

5. Given the reaction:



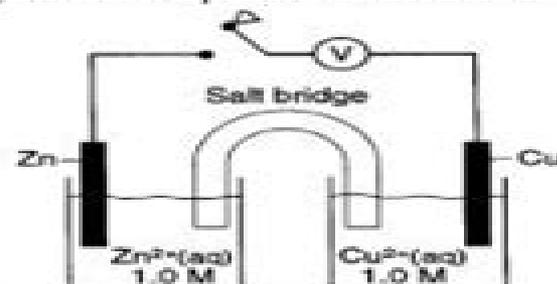
In which type of cell would this reaction most likely occur?

- A) a chemical cell, because it is exothermic
B) an electrolytic cell, because it is exothermic
C) a chemical cell, because it is endothermic
D) an electrolytic cell, because it is endothermic

6. Which half-reaction occurs at the negative electrode in an electrolytic cell in which an object is being plated with silver?

- A) $\text{Ag}^0 + 1e^- \rightarrow \text{Ag}^+$ B) $\text{Ag}^0 \rightarrow \text{Ag}^+ + 1e^-$
C) $\text{Ag}^+ + 1e^- \rightarrow \text{Ag}^0$ D) $\text{Ag}^+ \rightarrow \text{Ag}^0 + 1e^-$

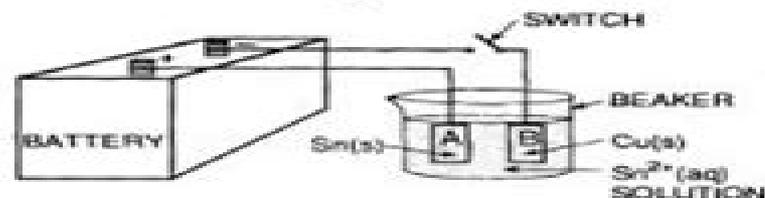
7. The diagram below represents an electrochemical cell.



What occurs when the switch is closed?

- A) Zn is reduced.
B) Cu is oxidized.
C) Electrons flow from Cu to Zn.
D) Electrons flow from Zn to Cu.

8. Base your answer to the following question on the diagram below of an electrolytic cell in which the electrodes are tin and copper.



In this electrolytic cell, electrode A is designated as the

- A) anode and is positive
B) anode and is negative
C) cathode and is positive
D) cathode and is negative

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