

**Atomic Orbitals**

Use with Chapter 5,  
Section 5.2

1. What is the shape of an s orbital?  
Spherical
2. What is the relationship between the size of an s orbital and the principal energy level in which it is found?  
\_\_\_\_\_
3. What is the shape of a p orbital? How many p orbitals are there in a sublevel?  
\_\_\_\_\_
4. How many electrons can each orbital hold?  
2
5. Look at the diagrams of the p orbitals. What do x, y, and z refer to?  
\_\_\_\_\_
6. How many d orbitals are there in a given sublevel? How many total electrons can the d orbitals in a sublevel hold?  
5 10
7. Which d orbitals have the same shape?  
 $x^2 - y^2$
8. What point in each diagram represents an atom's nucleus?  
\_\_\_\_\_
9. How likely is it that an electron occupying a p or a d orbital would be found very near an atom's nucleus? What part of the diagram supports your conclusion?  
\_\_\_\_\_

# Teaching Transparency 16 Answers

**Marilyn Jager Adams**



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