



COVALENT BONDING

Practice Problems

Do your best to solve each of the following problems.

SECTION 8.1 MOLECULAR COMPOUNDS

- Classify each of the following as either an ionic or a molecular compound.
 - NaCl
 - H_2O
 - HCl
 - KNO_3
 - CaCl_2
 - H_2SO_4
 - C_2H_6
 - CO_2

- Classify each of the following as either an ionic or a molecular compound. Write the name of the compound.
- Write the formula for each of the following molecular compounds. Write the name of the compound.
- Write the formula for each of the following ionic compounds. Write the name of the compound.

SECTION 8.2 THE NATURE OF COVALENT BONDING

- Draw a Lewis structure for hydrogen fluoride. $\text{H}-\text{F}$
- Draw a Lewis structure for carbon dioxide. $\text{O}=\text{C}=\text{O}$
- Draw a Lewis structure for methane. C
- Draw a Lewis structure for ethane. $\text{H}-\text{C}-\text{C}-\text{H}$
- Draw a Lewis structure for acetylene. $\text{H}-\text{C}\equiv\text{C}-\text{H}$
- Draw a Lewis structure for ammonia. N



SECTION 8.3 BONDING: THE ORBITALS

- Which orbital and sublevel are represented by the quantum numbers $n=2$, $l=1$, $m_l=0$, and $m_s=+\frac{1}{2}$?
- Which orbital and sublevel are represented by the quantum numbers $n=3$, $l=2$, $m_l=0$, and $m_s=+\frac{1}{2}$?

Pearson Chapter 8 Covalent Bonding Answers

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Table of Contents Pearson Chapter 8 Covalent Bonding Answers

1. Understanding the eBook Pearson Chapter 8 Covalent Bonding Answers
 - The Rise of Digital Reading Pearson Chapter 8 Covalent Bonding Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Pearson Chapter 8 Covalent Bonding Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pearson Chapter 8 Covalent Bonding Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pearson Chapter 8 Covalent Bonding Answers
 - Personalized Recommendations
 - Pearson Chapter 8 Covalent Bonding Answers User Reviews and Ratings
 - Pearson Chapter 8 Covalent Bonding Answers and Bestseller Lists
5. Accessing Pearson Chapter 8 Covalent Bonding Answers Free and Paid eBooks
 - Pearson Chapter 8 Covalent Bonding Answers Public Domain eBooks
 - Pearson Chapter 8 Covalent Bonding Answers eBook Subscription Services
 - Pearson Chapter 8 Covalent Bonding Answers Budget-Friendly Options

6. Navigating Pearson Chapter 8 Covalent Bonding Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Pearson Chapter 8 Covalent Bonding Answers Compatibility with Devices
 - Pearson Chapter 8 Covalent Bonding Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pearson Chapter 8 Covalent Bonding Answers
 - Highlighting and Note-Taking Pearson Chapter 8 Covalent Bonding Answers
 - Interactive Elements Pearson Chapter 8 Covalent Bonding Answers
8. Staying Engaged with Pearson Chapter 8 Covalent Bonding Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pearson Chapter 8 Covalent Bonding Answers
9. Balancing eBooks and Physical Books Pearson Chapter 8 Covalent Bonding Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pearson Chapter 8 Covalent Bonding Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Pearson Chapter 8 Covalent Bonding Answers
 - Setting Reading Goals Pearson Chapter 8 Covalent Bonding Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Pearson Chapter 8 Covalent Bonding Answers
 - Fact-Checking eBook Content of Pearson Chapter 8 Covalent Bonding Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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