

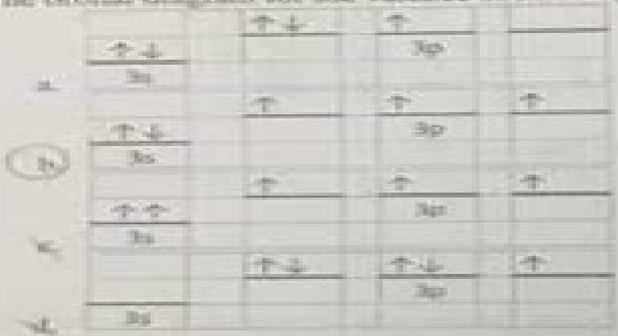
7.) Naturally-occurring copper exists as two isotopes: ^{63}Cu and ^{65}Cu , the masses of which are given below. Which of these isotopes, if either, is most abundant?

Isotope	Mass (amu)
^{63}Cu	62.929598
^{65}Cu	64.927730

- 35%
- a. More information is needed
 b. Neither, they are nearly-equal abundance
 c. ^{63}Cu is more abundant
 d. ^{65}Cu is more abundant

We can find the answer given the information and they would be abundant

8.) The orbital diagram for the valence shell of a phosphorous atom in its ground state is:



Section II – Compounds, Reactions, Mole Concept, Solutions, and Stoichiometry

9.) How many moles of xenon tetrafluoride correspond to 1.036 g of this compound?

- a. 2.5 mmol
 b. 1.5 mmol
 c. 500. mmol
 d. 5.0 mmol
 e. None of these

$1.036 \text{ g } \text{XeF}_4 \times \frac{1 \text{ mol}}{354.5 \text{ g}} = 0.0029 \text{ mol} \times 1000 \frac{\text{mmol}}{\text{mol}}$

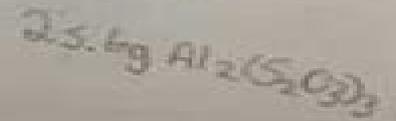
10.) How many iodine atoms are present in 39.0 g of $\text{Ca}(\text{IO}_3)_2$?

- a. 6.002×10^{23}
 b. 3.011×10^{23}
 c. 1.20×10^{24}
 d. 1.20×10^{22}
 e. None of these

$\text{Ca}(\text{IO}_3)_2 \quad 389.88 \text{ g mol}^{-1}$



$39.0 \text{ g} \times \frac{1 \text{ mol}}{389.88} \times \frac{2 \text{ mol I}}{1 \text{ mol}} = \text{mol I}$
 $= 0.200$



$4.0 \times 10^{-1} \text{ mol} \times \frac{1 \text{ mol}}{1 \text{ mol}}$

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