

Chapter 2 Chemistry of Life

- Matter anything that takes up space
- An Element is one of the basic building blocks of matter and cannot be broken down by chemical means.
- It is even more surprising that that over 90% of the human body is composed of just four elements.
 - Carbon
 - Nitrogen
 - Oxygen
 - Hydrogen
- Atoms are the simplest unit of element that still retains the chemical and physical properties of an element.
 - There are the best-known subatomic particles are positively charged protons
 - Uncharged: Neutrons
 - Negatively charged: Electrons
 - Protons and Neutrons are located within the nucleus of an atom and electrons move about the nucleus
 - The circle around the nucleus of an atom represents an electron shell, which represents the average location of electrons
- The atomic number represents the number of protons in the nucleus of an element
- Mass number of an atom represents the sum of protons and neutrons in the nucleus
- The atomic mass (the number below the atomic symbol on the periodic table) is the average of the AMU for all the isotopes of that atom. To determine the number of neutrons, subtract the number of protons from the atomic mass and take the closest whole number.
- Isotopes of an atom have the same number of protons but different numbers
- Atom with more than one shell are most stable when the outer shell is also called shell contains 8 electrons.
- Ions are particles that either carry either positive or negative charge.
- The attraction between oppositely charged sodium ions and chloride ions form an ionic bond.
- Atoms share in covalent bonds
- A hydrogen bond is the attraction of a slightly positive covalently bonded hydrogen to a slightly negative atom in the vicinity
- Water has a high heat capacity
- Water has a high heat of evaporation
- Water molecules are cohesive and adhesive
- Cohesion refers to the ability of water molecules to cling to each other
- Adhesion refers to the ability of water molecule to cling to a surface
- Lemon juice, Vinegar, Tomatoes, and coffee are all acidic solutions
- Acids are substances that dissociates in water releasing hydrogen ions. The acidity of a substance depends on how fully it dissociates in water

Biology Chapter 2 The Chemistry Of Life Study Guide

Sylvia S. Mader



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Table of Contents Biology Chapter 2 The Chemistry Of Life Study Guide

1. Understanding the eBook Biology Chapter 2 The Chemistry Of Life Study Guide
 - The Rise of Digital Reading Biology Chapter 2 The Chemistry Of Life Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Biology Chapter 2 The Chemistry Of Life Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Biology Chapter 2 The Chemistry Of Life Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Biology Chapter 2 The Chemistry Of Life Study Guide
 - Personalized Recommendations
 - Biology Chapter 2 The Chemistry Of Life Study Guide User Reviews and Ratings
 - Biology Chapter 2 The Chemistry Of Life Study Guide and Bestseller Lists
5. Accessing Biology Chapter 2 The Chemistry Of Life Study Guide Free and Paid eBooks
 - Biology Chapter 2 The Chemistry Of Life Study Guide Public Domain eBooks
 - Biology Chapter 2 The Chemistry Of Life Study Guide eBook Subscription Services
 - Biology Chapter 2 The Chemistry Of Life Study Guide Budget-Friendly Options

6. Navigating Biology Chapter 2 The Chemistry Of Life Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Biology Chapter 2 The Chemistry Of Life Study Guide Compatibility with Devices
 - Biology Chapter 2 The Chemistry Of Life Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Biology Chapter 2 The Chemistry Of Life Study Guide
 - Highlighting and Note-Taking Biology Chapter 2 The Chemistry Of Life Study Guide
 - Interactive Elements Biology Chapter 2 The Chemistry Of Life Study Guide
8. Staying Engaged with Biology Chapter 2 The Chemistry Of Life Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Biology Chapter 2 The Chemistry Of Life Study Guide
9. Balancing eBooks and Physical Books Biology Chapter 2 The Chemistry Of Life Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Biology Chapter 2 The Chemistry Of Life Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Biology Chapter 2 The Chemistry Of Life Study Guide
 - Setting Reading Goals Biology Chapter 2 The Chemistry Of Life Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Biology Chapter 2 The Chemistry Of Life Study Guide
 - Fact-Checking eBook Content of Biology Chapter 2 The Chemistry Of Life Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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