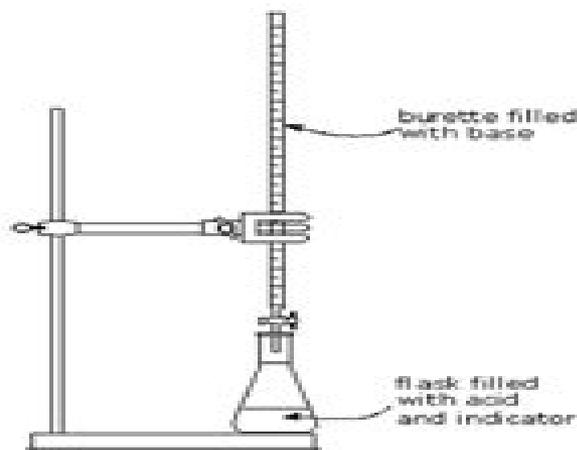


Titration Lab

Problem: To determine the concentration of the unknown acid or base.

Set-up:



Procedure:

1. Place 10 mL of HCl in the Erlenmeyer flask.
2. Place 5 drops of phenolphthalein in the Erlenmeyer flask with the HCl.
3. Fill the burette with NaOH using a funnel.
4. Place the burette in the retort stand to look like the set up in the diagram.
5. Place a piece of white paper underneath the Erlenmeyer flask.
6. Write down the initial volume of NaOH on the burette.
7. Slowly (drop by drop) add the NaOH to the Erlenmeyer flask, while swirling the liquid, until the solution stays a very light pink colour.
8. Write down the final volume of the NaOH on the burette.
9. Rinse out the Erlenmeyer flask and repeat the experiment at least 3 times.

Observations (4):

| | Trial 1 | Trial 2 | Trial 3 | Trial 4 |
|---------------------------------------|---------|---------|---------|---------|
| Initial (I) Volume of NaOH (mL) | 0.00 mL | 11.1 mL | 22.5 mL | 33.8 mL |
| Final (F) Volume of NaOH (mL) | 11.1 mL | 22.5 mL | 33.8 mL | 45.3 mL |
| Total Volume of NaOH used (mL) (I-F)* | 11.1 mL | 11.4 mL | 11.3 mL | 11.5 |

* the burette is labelled as full at 50 mL and goes down from there, so you will subtract the final value from the initial to get the total volume used.

Titration Lab Answers

Thomas Griffiths



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Table of Contents Titration Lab Answers

1. Understanding the eBook Titration Lab Answers
 - The Rise of Digital Reading Titration Lab Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Titration Lab Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Titration Lab Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Titration Lab Answers
 - Personalized Recommendations
 - Titration Lab Answers User Reviews and Ratings
 - Titration Lab Answers and Bestseller Lists
5. Accessing Titration Lab Answers Free and Paid eBooks
 - Titration Lab Answers Public Domain eBooks
 - Titration Lab Answers eBook Subscription Services
 - Titration Lab Answers Budget-Friendly Options
6. Navigating Titration Lab Answers eBook Formats

- ePub, PDF, MOBI, and More
 - Titration Lab Answers Compatibility with Devices
 - Titration Lab Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Titration Lab Answers
 - Highlighting and Note-Taking Titration Lab Answers
 - Interactive Elements Titration Lab Answers
 8. Staying Engaged with Titration Lab Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Titration Lab Answers
 9. Balancing eBooks and Physical Books Titration Lab Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Titration Lab Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Titration Lab Answers
 - Setting Reading Goals Titration Lab Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Titration Lab Answers
 - Fact-Checking eBook Content of Titration Lab Answers
 - Distinguishing Credible Sources
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 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
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