



Name:

Date:

Student Exploration: Phase Changes (4b)

Directions: Follow the instructions to go through the simulation. Respond to the questions and prompts in the orange boxes.

Vocabulary: altitude, boil, boiling point, freeze, freezing point, gas, liquid, melt, melting point, phase, solid

Activity B:

Temperature and molecular motion

Get the Gizmo ready:

- Click **Reset**, and select the **Micro view**.
- Set **Ice volume** to 0 cm^3 .
- Set **Add/remove heat energy** to 0 J/s .



Question: Why do phase changes occur?

1. **Compare:** Set the **Water temperature** to 0 $^{\circ}\text{C}$ and click **Play**. Observe the water molecules. Click **Reset**, set the **Water temperature** to 100 $^{\circ}\text{C}$, and click **Play** again.

What do you notice?

There is no difference.

2. **Observe:** Click **Reset**. The **mean molecular speed** of the water molecules is displayed below the container. Set the **Water temperature** to 0 $^{\circ}\text{C}$ and **Add/remove heat energy** to 400 J/s . Click **Play**.

A. How does the mean speed of the water molecules change as they are heated?

The water gets hotter (has more heat energy).

B. Does the mean molecular speed change as much as the temperature as the water heats up? Explain.

Yes, the more the temperature increases the more the molecules of the water speed up (until they hit the boiling point, where they cannot get hotter any longer.)

3. **Explain:** How is temperature related to the motions of molecules?

The higher the temperature, the faster the molecules move.

4. **Observe:** Click **Reset**. Set the **Water temperature** to 20 $^{\circ}\text{C}$ and the **Ice volume** to 50 cm^3 . Set **Add/remove heat energy** to 0 J/s . Click **Play**. How do the molecules in the liquid interact with the molecules in the solid?

Gizmo Phase Changes Answers

Lei Huang



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Table of Contents Gizmo Phase Changes Answers

1. Understanding the eBook Gizmo Phase Changes Answers
 - The Rise of Digital Reading Gizmo Phase Changes Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Gizmo Phase Changes Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gizmo Phase Changes Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gizmo Phase Changes Answers
 - Personalized Recommendations
 - Gizmo Phase Changes Answers User Reviews and Ratings
 - Gizmo Phase Changes Answers and Bestseller Lists
5. Accessing Gizmo Phase Changes Answers Free and Paid eBooks
 - Gizmo Phase Changes Answers Public Domain eBooks
 - Gizmo Phase Changes Answers eBook Subscription Services
 - Gizmo Phase Changes Answers Budget-Friendly Options

6. Navigating Gizmo Phase Changes Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Gizmo Phase Changes Answers Compatibility with Devices
 - Gizmo Phase Changes Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gizmo Phase Changes Answers
 - Highlighting and Note-Taking Gizmo Phase Changes Answers
 - Interactive Elements Gizmo Phase Changes Answers
8. Staying Engaged with Gizmo Phase Changes Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gizmo Phase Changes Answers
9. Balancing eBooks and Physical Books Gizmo Phase Changes Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gizmo Phase Changes Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Gizmo Phase Changes Answers
 - Setting Reading Goals Gizmo Phase Changes Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gizmo Phase Changes Answers
 - Fact-Checking eBook Content of Gizmo Phase Changes Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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