

CHIMICA GENERALE

Ogni atomo/elemento è caratterizzato da una simbolo

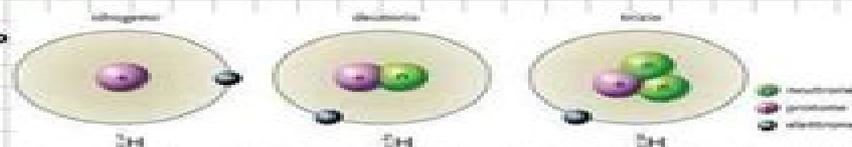


- il numero di massa è dato dalla somma di protoni + neutroni (nucleoni)
- il numero atomico corrisponde al numero dei protoni (uguale al numero di e⁻)

un nucleo con particolari valori di Z e A → **NUCLIDE**

GLI ISOTOP:

Atomi di uno stesso elemento che possono differire per il numero di neutroni. I diversi isotopi di un elemento hanno uguali proprietà chimiche, ma diverse proprietà fisiche.



→ **CATIONI:** quando ad un atomo vengono sottratti uno o più elettroni, si formano ioni carichi positivamente.

→ **ANIONI:** quando l'atomo assume uno o più elettroni si formano ioni carichi negativamente.

- contiene 7 periodi (orizzontali)
- 18 gruppi (verticali) → nei gruppi gli elementi hanno caratteristiche chimiche simili

TAVOLA DEGLI ELEMENTI:

essi sono organizzati per massa atomica crescente.

gruppi	1A	2A	3A - 10A										11A	12A	13A	14A	15A	16A	17A	18A																
1	H	He																																		
2	Li	Be											B	C	N	O	F	Ne																		
3	Na	Mg	Al	Si	P	S	Cl	Ar	K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr										
4	K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr	Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe
5	Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe	Cs	Ba	La	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	
6	Cs	Ba	La	Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn				
7	Fr	Ra	Ac	Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Mn	Nv	Og	118	119	120	121	122	123	124	125	126	127	128	129	130	131	132	133	134	135	136

■ orbitale s
■ orbitale p
■ orbitale d
■ orbitale f

- **GRUPPO 1:** metalli alcalini (Li, Na, K, Rb, Cs)
- **GRUPPO 2:** metalli alcalino-terrosi (Be, Mg, Ca, Sr, Ba)
- essi reagiscono con l'acqua producendo idrogeno e calore es il potassio reagisce tanto da fiammare l'idrogeno.
- **GRUPPO 18:** gas nobili, essi si combinano con pochissimi altri elementi. Sono tutti gas incolori e inodori. Sono poco reattivi.
- **GRUPPO 17:** alogeni. Si presentano con molecole biatomiche.

- **I METALLI:** sono solidi, dotati di lucentezza, malleabilità e duttilità; essi sono buoni conduttori di elettricità e calore.

- **I NON METALLI:** sono gas o solidi che non hanno caratteristiche metalliche.

- **SEMIMETALLI O METALLOIDI:** hanno caratteristiche intermedie tra i metalli e i non metalli.

Appunti Di Chimica Generale Unife

Mike Jess



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