

AP Chemistry Chapter 4 Practice Exam

- D 1. The Claus reactions, shown below, are used to generate elemental sulfur from hydrogen sulfide.



How much sulfur (in grams) is produced from 21.5 grams of O_2 ?

A) 97.0 g

B) 14.4 g

C) 21.5 g

D) 43.1 g

E) none of these

$$\frac{21.5 \text{ g O}_2}{32.00 \text{ g O}_2} \times \frac{1 \text{ mol O}_2}{3 \text{ mol O}_2} \times \frac{2 \text{ SO}_2}{1 \text{ SO}_2} \times \frac{3 \text{ S}}{1 \text{ SO}_2} \times \frac{32.06 \text{ g S}}{1 \text{ mol S}} = 43.1 \text{ g S}$$

- B 2. Iron is produced from its ore by the reactions:



How many moles of $\text{O}_2(g)$ are needed to produce 4.6 moles of $\text{Fe}(s)$?

A) 2.3 mole O_2

B) 3.5 mole O_2

C) 4.6 mole O_2

D) 6.9 mole O_2

E) 13.8 mole O_2

$$\frac{4.6 \text{ mol Fe}(s)}{2 \text{ Fe}} \times \frac{3 \text{ CO}}{2 \text{ CO}} \times \frac{1 \text{ O}_2}{2 \text{ CO}} = 3.45 \text{ mol O}_2$$

- B 3. Suppose the reaction $\text{Ca}_3(\text{PO}_4)_2 + 3\text{H}_2\text{SO}_4 \rightarrow 3\text{CaSO}_4 + 2\text{H}_3\text{PO}_4$ is carried out starting with 153 g of $\text{Ca}_3(\text{PO}_4)_2$

and 76.8 g of H_2SO_4 . How much phosphoric acid will be produced?

A) 76.7 g

B) 31.1 g

C) 229.8 g

D) 115.1 g

E) 96.7 g

$$\frac{153 \text{ g Ca}_3(\text{PO}_4)_2}{310 \text{ g Ca}_3(\text{PO}_4)_2} \times \frac{1 \text{ mol Ca}_3(\text{PO}_4)_2}{1 \text{ mol Ca}_3(\text{PO}_4)_2} \times \frac{2 \text{ H}_3\text{PO}_4}{1 \text{ Ca}_3(\text{PO}_4)_2} = 0.97 \text{ mol H}_3\text{PO}_4$$

$$\frac{76.8 \text{ g H}_2\text{SO}_4}{98 \text{ g H}_2\text{SO}_4} \times \frac{1 \text{ mol H}_2\text{SO}_4}{3 \text{ mol H}_2\text{SO}_4} \times \frac{2 \text{ H}_3\text{PO}_4}{1 \text{ mol H}_2\text{SO}_4} = 0.52 \text{ mol H}_3\text{PO}_4$$

- B 4. SO_2 reacts with H_2S as follows:



When 7.50 g of H_2S reacts with 12.75 g of SO_2 , which statement applies?

A) 6.38 g of sulfur are formed.

B) 10.6 g of sulfur are formed.

C) 0.0216 moles of H_2S remain.

D) 1.13 g of H_2S remain.

E) SO_2 is the limiting reagent.

$$\frac{7.50 \text{ g H}_2\text{S}}{34.03 \text{ g H}_2\text{S}} \times \frac{1 \text{ mol H}_2\text{S}}{2 \text{ mol H}_2\text{S}} \times \frac{3 \text{ S}}{1 \text{ mol H}_2\text{S}} \times \frac{32 \text{ g S}}{1 \text{ mol S}} = 10.89$$

$$\frac{12.75 \text{ g SO}_2}{64 \text{ g SO}_2} \times \frac{1 \text{ mol SO}_2}{1 \text{ mol SO}_2} \times \frac{3 \text{ S}}{1 \text{ mol SO}_2} \times \frac{32 \text{ g S}}{1 \text{ mol S}} = 19.13$$

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Table of Contents Ap Chem Chapter 1 Practice Test

1. Understanding the eBook Ap Chem Chapter 1 Practice Test
 - The Rise of Digital Reading Ap Chem Chapter 1 Practice Test
 - Advantages of eBooks Over Traditional Books
2. Identifying Ap Chem Chapter 1 Practice Test
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ap Chem Chapter 1 Practice Test
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ap Chem Chapter 1 Practice Test
 - Personalized Recommendations
 - Ap Chem Chapter 1 Practice Test User Reviews and Ratings
 - Ap Chem Chapter 1 Practice Test and Bestseller Lists
5. Accessing Ap Chem Chapter 1 Practice Test Free and Paid eBooks

- Ap Chem Chapter 1 Practice Test Public Domain eBooks
 - Ap Chem Chapter 1 Practice Test eBook Subscription Services
 - Ap Chem Chapter 1 Practice Test Budget-Friendly Options
6. Navigating Ap Chem Chapter 1 Practice Test eBook Formats
- ePub, PDF, MOBI, and More
 - Ap Chem Chapter 1 Practice Test Compatibility with Devices
 - Ap Chem Chapter 1 Practice Test Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Ap Chem Chapter 1 Practice Test
 - Highlighting and Note-Taking Ap Chem Chapter 1 Practice Test
 - Interactive Elements Ap Chem Chapter 1 Practice Test
8. Staying Engaged with Ap Chem Chapter 1 Practice Test
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ap Chem Chapter 1 Practice Test
9. Balancing eBooks and Physical Books Ap Chem Chapter 1 Practice Test
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ap Chem Chapter 1 Practice Test
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Ap Chem Chapter 1 Practice Test
- Setting Reading Goals Ap Chem Chapter 1 Practice Test
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ap Chem Chapter 1 Practice Test
- Fact-Checking eBook Content of Ap Chem Chapter 1 Practice Test
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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