

KINETICS Practice Problems and Solutions

Name: _____
 Period: _____ Date: _____

AP Chemistry
 Dr. Mandes

The following questions represent potential types of quiz questions. Please answer each question completely and thoroughly. The solutions will be posted on-line on Monday.

5. Please do #18 in chapter 12 of your text.
- This increases the $[H_2]$, which will increase the rate, but has no effect on k
 - Due to Arrhenius equation, changing temperature changes the value of k .
 - Catalyst function by lowering the activation energy, so due to Arrhenius equation, changing the activation energy changes the value of k .

6. Consider the reaction: $P_4 + 6 H_2 \rightarrow 4 PH_3$. A rate study of this reaction was conducted at 298 K. The data that were obtained are shown in the table.

$[P_4]$, mol/L	$[H_2]$, mol/L	Initial Rate, mol/L · s
0.0110	0.0075	3.20×10^{-4}
0.0110	0.0150	6.40×10^{-4}
0.0220	0.0150	6.39×10^{-4}

- What is the order with respect to: P_4 0
 H_2 1
- Write the rate law for this reaction. $rate = k[H_2]$
- Determine the value and units of the rate constant, k . **plug and chug using the rate law & data from exp't 1 and solving for k , we get $k = 0.0427 s^{-1}$**

7. Consider the reaction: $SO_2 + O_3 \rightarrow SO_3 + O_2$. A rate study of this reaction was conducted at 298 K. The data that were obtained are shown in the table.

$[SO_2]$, mol/L	$[O_3]$, mol/L	Initial Rate, mol/L · s
0.25	0.40	0.118
0.25	0.20	0.118
0.75	0.20	1.062

- What is the order with respect to: SO_2 2
 O_3 0
- Write the rate law for this reaction. $rate = k[SO_2]^2[O_3]^0$
- Determine the value and units of the rate constant, k . **plug and chug using the rate law & data from exp't 1 and solving for k , we get $k = 2.36 molL^{-2} s^{-1}$**

8. Consider the following mechanism.
- $$A_2 + B_2 \rightarrow R + C \quad (\text{slow})$$
- $$A_2 + R \rightarrow C \quad (\text{fast})$$

- Write the overall balanced chemical equation. $2 A_2 + B_2 \rightarrow 2 C$
- Identify any intermediates within the mechanism. **R**
- What is the order with respect to each reactant? **A_2 : 1st; B_2 : 1st**

Solutions Chemical Kinetics

Paul L. Houston



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Chemical Kinetics Kenneth Antonio Connors,1990 **Chemical Kinetics The Study of Reaction Rates in Solution** Kenneth A Connors This chemical kinetics book blends physical theory phenomenology and empiricism to provide a guide to the experimental practice and interpretation of reaction kinetics in solution It is suitable for courses in chemical kinetics at the graduate and advanced undergraduate levels This book will appeal to students in physical organic chemistry physical inorganic chemistry biophysical chemistry biochemistry pharmaceutical chemistry and water chemistry all fields concerned with the rates of chemical reactions in the solution phase Solutions Manual: Sm Chemical Kinetics and React Dyn Paul L. Houston,2001 **Catalog of National Bureau of Standards Publications, 1966-1976: Citations and abstracts** United States. National Bureau of Standards. Technical Information and Publications Division,1978 Publications United States. National Bureau of Standards,1978 **Catalog of National Bureau of Standards Publications, 1966-1976** United States. National Bureau of Standards,1978 **Theoretical Chemistry from the Standpoint of Avogadro's Rule & Thermodynamics** Walther Nernst,1923 **The Industrial Arts Index** ,1930 **Scientific and Technical Aerospace Reports** ,1966 **Publications of the National Bureau of Standards ... Catalog** United States. National Bureau of Standards,1978 **American Druggist** ,1923 **Reaction Kinetics** Keith J. Laidler,2013-10-22 **Reaction Kinetics Volume II Reactions in Solution** deals with the kinetics of reactions in solution and discusses the basic principles and theories of kinetics including a brief description of homogeneous gas reactions This book is divided into two chapters The first chapter focuses on the general principles of reactions in solution that includes reactions between ions and involving dipoles influence of pressure on rates in solution substituent effects and homogeneous catalysis in solution Chapter 2 primarily deals with general features of reactions in solution emphasizing the relationship between the results of a kinetic investigation and actual reaction mechanism This volume is intended for undergraduate students of chemistry who have not previously studied chemical kinetics This book is also useful to more advanced students in other fields such as biology and physics who wish to have a general knowledge of the subject **Chemistry** John Christian Bilar,1989 **College of Engineering Catalogue** Carnegie Institute of Technology. College of Engineering and Science,1927 **Classification for Works on Pure and Applied Science in the Science Library, the Science Museum** Science Museum (Great Britain). Library,1921 **International Journal of Chemical Kinetics** ,1993 Topics covered in this publication include quantitative relationships between molecular structure and chemical activity organic inorganic chemistry biochemical kinetics and reaction mechanisms Surface kinetics are also explored *Physical Review* ,1897 Vols for 1903 include Proceedings of the American Physical Society **Biennial Catalog** New Mexico Institute of Mining and Technology,New Mexico. Institute of Mining and Technology,1893 *Industrial Arts Index* ,1930 **Calendar of the University of Sydney** University of Sydney,1918 Science Progress ,1916

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