

Subject:- Chemistry

# Chemical Thermodynamics and Energetics

## Practice Problems-II

Q1) Three moles of an ideal gas are expanded isothermally from a volume of 300 cm<sup>3</sup> to 2.5 L at 300 K against a pressure of 1.9 atm. Calculate the work done in L atm and joules.

Soln:-  $V_1 = 300 \text{ c.c.} = 0.3 \text{ L}$ ,  $V_2 = 2.5 \text{ L}$ ,  $T = 300 \text{ K}$ ,  $P_{ex} = 1.9 \text{ atm}$

$$W = -P_{ex}\Delta V = -P_{ex}(V_2 - V_1) \\ = -1.9(2.5 - 0.3) = -1.9 \times 2.2$$

$$\therefore W = -4.18 \text{ L atm}$$

$$\text{and } W = -4.18 \times 101.3 = -423.4 \text{ J}$$

Q2) One mole of an ideal gas is compressed from 500 cm<sup>3</sup> against a constant pressure of  $1.216 \times 10^5 \text{ Pa}$ . The work involved in the process is 36.50 J. Calculate the final volume.

Soln:-  $V_1 = 500 \text{ c.c.} = 0.5 \text{ dm}^3$ ,

$$P_{ex} = 1.216 \times 10^5 \text{ Pa} = 1.216 \times 10^2 \text{ kPa}, W = 36.50 \text{ J}$$

$$W = -P_{ex}\Delta V$$

# Problems In Chemical Thermodynamics With Solutions

**KJ Lindholm-Leary**



## **Problems In Chemical Thermodynamics With Solutions:**

Problems in Chemical Thermodynamics with Solutions Maka Aleksishvili, Shota Sidamonidze, 2002 The methods of chemical thermodynamics are effectively used in many fields of science and technology Mastering these methods and their use in practice requires profound comprehension of the theoretical questions and acquisition of certain calculating skills This book is useful to undergraduate and graduate students in chemistry as well as chemical thermal and refrigerating technology it will also benefit specialists in all other fields who are interested in using these powerful methods in their practical activities

*Chemical Thermodynamics* M L McGlashan, 2007-10-31 Specialist Periodical Reports provide systematic and detailed review coverage of progress in the major areas of chemical research Written by experts in their specialist fields the series creates a unique service for the active research chemist supplying regular critical in depth accounts of progress in particular areas of chemistry For over 80 years the Royal Society of Chemistry and its predecessor the Chemical Society have been publishing reports charting developments in chemistry which originally took the form of Annual Reports However by 1967 the whole spectrum of chemistry could no longer be contained within one volume and the series Specialist Periodical Reports was born The Annual Reports themselves still existed but were divided into two and subsequently three volumes covering Inorganic Organic and Physical Chemistry For more general coverage of the highlights in chemistry they remain a must Since that time the SPR series has altered according to the fluctuating degree of activity in various fields of chemistry Some titles have remained unchanged while others have altered their emphasis along with their titles some have been combined under a new name whereas others have had to be discontinued

*Practical Chemical Thermodynamics for Geoscientists* Bruce Fegley, Rose Osborne, 2013 Presents brief historical summaries and biographies of key thermodynamics scientists alongside the fundamentals they were responsible for

*Basic Chemical Thermodynamics* Eric Brian Smith, 2004 This widely acclaimed text now in its fifth edition and translated into many languages continues to present a clear simple and concise introduction to chemical thermodynamics An examination of equilibrium in the everyday world of mechanical objects provides the starting point for an accessible account of the factors that determine equilibrium in chemical systems This straightforward approach leads students to a thorough understanding of the basic principles of thermodynamics which are then applied to a wide range of physico chemical systems The book also discusses the problems of non ideal solutions and the concept of activity and provides an introduction to the molecular basis of thermodynamics Over five editions the views of teachers of the subject and their students have been incorporated The result is a little more rigour in specifying the dimensions within logarithmic expressions the addition of more worked examples and the inclusion of a simple treatment of the molecular basis of thermodynamics Students on courses in thermodynamics will continue to find this popular book an excellent introductory text

*Chemical Thermodynamics* J. Rouquerol, R. Sabbah, 2013-10-22 Chemical Thermodynamics 4 presents the application of experimental methods of chemical thermodynamics This book discusses the three properties of

biological molecules namely colossal dimension exclusive orderliness and capability to be in different states or conformations depending on conditions Organized into eight chapters this book begins with an overview of the trends in thermochemistry that involve complex reaction systems and product mixtures This text then discusses the problems relating to the standard state of solids and illustrates the utilization of enthalpy of mixing data Other chapters consider the available heat capacity results in the liquid gas This book discusses as well the high temperature measurement of thermodynamic data for substances of metallurgical interest The final chapter deals with the important advances in the experimental methods of heat capacity measurements including laser flash calorimetry and the high resolution heat capacity calorimeter This book is a valuable resource for chemists physical chemists thermochemists thermophysicists nuclear engineers and research workers

*Bulletin of Chemical Thermodynamics*, 1987      **Concepts And Problems In Physical Chemistry** P.S. Raghavan, 1997

Contents Introduction Atoms Molecules and Formulas Chemical Equations and Stoichiometry Aqueous Reactions and Solution Stoichiometry Gases Intermolecular Forces Liquids and Solids Atoms Structure and the Periodic Table Chemical Bonding Chemical Thermodynamics Solutions Chemical Kinetics Chemical Equilibrium Acids and Bases Ionic Equilibria I Ionic Equilibria II Redox Reactions Electrochemistry Nuclear Chemistry      **Thermodynamics and the Free Energy of Chemical Substances** Gilbert Newton Lewis, Merle Randall, 1923 The scope of thermodynamics Definitions the concept of equilibrium Conventions and mathematical methods Solutions The first law of thermodynamics and the concept of energy The fugacity Application of the second law to solutions The perfect solution The laws of the dilute solution Systems involving variables other than pressure temperature and composition A useful function called the activity and its application to solutions Change of activity with the temperature and the calculation of activity from freezing points The standard change of free energy the equilibrium constant Solutions of electrolytes The activity of strong electrolytes The activity of electrolytes from freezing point data and tables of activity coefficients Activity coefficient in mixed electrolytes the principle of the ionic strength the activity of individual ions The galvanic cell Single potentials standard electrode potentials of the elements The third law of thermodynamics The entropy of monatomic gases and a table of atomic entropies Introduction to systematic free energy calculations the free energy of elementary hydrogen and metallic hydrides Oxygen and its compounds with hydrogen and with some metals Chlorine and its compounds Bromine and its compounds Iodine and its compounds Nitrogen compounds Carbon and some of its compounds Compounds of carbon and nitrogen Table of free energies and examples illustrating its use Conversion table for mol fractions mol ratios and molities Some useful numerical factors Coefficients employed in converting activity equilibrium constant and free energy from one temperature to another Publications by the authrs pertaining to thermodynamics      **A System of Physical Chemistry: Thermodynamics** William Cudmore McCullagh Lewis, 1922      *The Chemical News and Journal of Physical Science*, 1913      *The Chemical News and Journal of Industrial Science* William Crookes, James H. Gardiner, Gerald Druce, H. W. Blood-Ryan, 1911      *The Montreal Medical Journal* George

Edgeworth Fenwick, Thomas George Roddick, George Ross, 1906 Chemical Addresses Delivered at the Second Decennial Celebration of Clark University Clark University (Worcester, Mass.), 1911 Catalogue Johns Hopkins University. College for Teachers, 1930 *Catalog Issue* Pratt Institute. School of Science and Technology, 1926 *Science*, 1926 **An Introduction to the principles of physical chemistry from the standpoint of modern atomistics and thermodynamics** Edward Wight Washburn, 1915 *The Chemical News*, 1911 *Chemical News and Journal of Physical Science*, 1902 *Johns Hopkins University Circulars* Johns Hopkins University, 1928

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