

## 11-2 Practice Problems

1. Determine the mass of lithium hydroxide produced when 0.38 g of lithium nitride reacts with water according to the following equation:



$$\frac{0.38\text{g Li}_3\text{N}}{1} \times \frac{1\text{mol Li}_3\text{N}}{35\text{g Li}_3\text{N}} \times \frac{3\text{mol LiOH}}{1\text{mol Li}_3\text{N}} \times \frac{24\text{g LiOH}}{1\text{mol LiOH}} = \boxed{0.78\text{g LiOH}}$$

2. What mass of sodium chloride is produced when chlorine reacts with 0.29 g of sodium iodide?



$$\frac{0.29\text{g NaI}}{1} \times \frac{1\text{mol}}{150} \times \frac{2\text{mol NaCl}}{2\text{mol NaI}} \times \frac{58\text{g}}{1\text{mol}} = \boxed{0.112\text{g NaCl}}$$

3. Determine the mass of carbon dioxide produced when 0.85 g of butane reacts with oxygen according to the following equation:



$$\frac{0.85\text{g C}_4\text{H}_{10}}{1} \times \frac{1\text{mol}}{58\text{g}} \times \frac{8\text{mol CO}_2}{2\text{mol C}_4\text{H}_{10}} \times \frac{44\text{g}}{1\text{mol}} = \boxed{2.58\text{g CO}_2}$$

4. Determine the mass of antimony produced when 0.46 g of antimony(III) oxide reacts with carbon according to the following equation:



$$\frac{0.46\text{g Sb}_2\text{O}_3}{1} \times \frac{1\text{mol}}{291.52\text{g}} \times \frac{2\text{mol Sb}}{1\text{mol Sb}_2\text{O}_3} \times \frac{121.76\text{g}}{1\text{mol}} = \boxed{0.38\text{g Sb}}$$

5. What mass of hydrogen peroxide ( $\text{H}_2\text{O}_2$ ) must decompose to produce 0.77 g of water?



$$\frac{0.77\text{g H}_2\text{O}}{1} \times \frac{1\text{mol}}{18\text{g}} \times \frac{2\text{H}_2\text{O}_2}{2\text{H}_2\text{O}} \times \frac{34\text{g}}{1\text{mol}} = \boxed{1.45\text{g H}_2\text{O}_2}$$

6. What mass of carbon monoxide must react with oxygen to produce 0.69 g of carbon dioxide?



$$\frac{0.69\text{g CO}_2}{1} \times \frac{1\text{mol}}{44\text{g}} \times \frac{2\text{mol CO}}{2\text{mol CO}_2} \times \frac{28\text{g}}{1\text{mol}} = \boxed{0.44\text{g CO}}$$

7. Determine the mass of sodium nitrate produced when 0.73 g of nickel(II) nitrate reacts with sodium hydroxide according to the following equation:



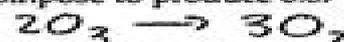
$$\frac{0.73\text{g Ni}(\text{NO}_3)_2}{1} \times \frac{1\text{mol}}{185\text{g}} \times \frac{2\text{mol NaNO}_3}{1\text{mol Ni}(\text{NO}_3)_2} \times \frac{85\text{g}}{1\text{mol}} = \boxed{0.68\text{g NaNO}_3}$$

8. Determine the mass of calcium hydroxide produced when calcium carbide reacts with 0.64 g of water according to the following equation:



$$\frac{0.64\text{g H}_2\text{O}}{1} \times \frac{1\text{mol}}{18\text{g}} \times \frac{1\text{mol Ca}(\text{OH})_2}{2\text{mol H}_2\text{O}} \times \frac{74\text{g}}{1\text{mol}} = \boxed{1.32\text{g Ca}(\text{OH})_2}$$

9. How many grams of ozone ( $\text{O}_3$ ) must decompose to produce 0.87 g of oxygen?



$$\frac{0.87\text{g O}_2}{1} \times \frac{1\text{mol}}{32\text{g O}_2} \times \frac{2\text{mol}}{3\text{mol}} \times \frac{48\text{g O}_3}{1\text{mol}} = \boxed{0.87\text{g O}_3}$$

10. Find the mass of sugar ( $\text{C}_6\text{H}_{12}\text{O}_6$ ) required to produce 1.82 L of carbon dioxide gas at STP from the reaction described by the following equation:



11. How many liters of oxygen are necessary for the combustion of 425 g of sulfur, assuming that the reaction occurs at STP? The balanced equation is  $\text{S} + \text{O}_2 \rightarrow \text{SO}_2$ .

12. Find the mass of benzene ( $\text{C}_6\text{H}_6$ ) required to produce 2.66 L of carbon dioxide gas at STP from the reaction described by the following equation:



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**JS Bruner**



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### **Table of Contents Prentice Hall Chemistry Workbook Answers Chapter 10**

1. Understanding the eBook Prentice Hall Chemistry Workbook Answers Chapter 10
  - The Rise of Digital Reading Prentice Hall Chemistry Workbook Answers Chapter 10
  - Advantages of eBooks Over Traditional Books
2. Identifying Prentice Hall Chemistry Workbook Answers Chapter 10
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Prentice Hall Chemistry Workbook Answers Chapter 10
  - User-Friendly Interface
4. Exploring eBook Recommendations from Prentice Hall Chemistry Workbook Answers Chapter 10

- Personalized Recommendations
  - Prentice Hall Chemistry Workbook Answers Chapter 10 User Reviews and Ratings
  - Prentice Hall Chemistry Workbook Answers Chapter 10 and Bestseller Lists
5. Accessing Prentice Hall Chemistry Workbook Answers Chapter 10 Free and Paid eBooks
    - Prentice Hall Chemistry Workbook Answers Chapter 10 Public Domain eBooks
    - Prentice Hall Chemistry Workbook Answers Chapter 10 eBook Subscription Services
    - Prentice Hall Chemistry Workbook Answers Chapter 10 Budget-Friendly Options
  6. Navigating Prentice Hall Chemistry Workbook Answers Chapter 10 eBook Formats
    - ePub, PDF, MOBI, and More
    - Prentice Hall Chemistry Workbook Answers Chapter 10 Compatibility with Devices
    - Prentice Hall Chemistry Workbook Answers Chapter 10 Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Prentice Hall Chemistry Workbook Answers Chapter 10
    - Highlighting and Note-Taking Prentice Hall Chemistry Workbook Answers Chapter 10
    - Interactive Elements Prentice Hall Chemistry Workbook Answers Chapter 10
  8. Staying Engaged with Prentice Hall Chemistry Workbook Answers Chapter 10
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Prentice Hall Chemistry Workbook Answers Chapter 10
  9. Balancing eBooks and Physical Books Prentice Hall Chemistry Workbook Answers Chapter 10
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Prentice Hall Chemistry Workbook Answers Chapter 10
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Prentice Hall Chemistry Workbook Answers Chapter 10
    - Setting Reading Goals Prentice Hall Chemistry Workbook Answers Chapter 10
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Prentice Hall Chemistry Workbook Answers Chapter 10

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  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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