



# Practical Object Oriented Design

**Léo Pomar**



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web a similar unit of concentration is molality m which is defined as the number of moles of solute per kilogram of solvent not per liter of solution molality  $\frac{\text{moles solute}}{\text{kilograms solvent}}$  mathematical manipulation of molality is the

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web 1 molarity  $M = \frac{\text{mass}}{\text{molar mass} \times \text{volume}}$  0.4500 l 0.825 g 141.9579  $\text{g mol}^{-1}$  0.0129 m 2 molality  $m = \frac{\text{mass}}{\text{molar mass} \times \text{mass solvent}}$  0.825 g 141.9579  $\text{g mol}^{-1}$  0.00581158 mol 0.00581158 mol 0.4500 kg 0.0129 m 3 mole fraction  $x = \frac{\text{mass}}{\text{molar mass} \times \text{total mass}}$  2 hpo 4 0.825 g 141.9579  $\text{g mol}^{-1}$  0.00581158 mol h 2 o 450.0 g 18.015  $\text{g mol}^{-1}$  24.97918401 mol

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$$\frac{\text{mass of a component}}{\text{mass of component in solution total mass of solution}} \times 100$$
 end array

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$\tanh x$   $2 \ 3 \ x$   $\operatorname{sech} 2 \ x$   $2 \ 3 \ x$   $2 \ x$   $3$

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