

## CHAPTER 15 STUDY GUIDE

## Energy and Chemical Change

### Section 15.1 Energy

In your textbook, read about the nature of energy.

In the space at the left, write *true* if the statement is true; if the statement is false, change the italicized word or phrase to make it true.

- \_\_\_\_\_ 1. *Energy* is the ability to do work or produce heat.
- \_\_\_\_\_ 2. The law of conservation of energy states that energy *can be* created and destroyed.
- \_\_\_\_\_ 3. Chemical potential energy is energy stored in a substance because of its *composition*.
- \_\_\_\_\_ 4. *Heat* is a form of energy that flows from a warmer object to a cooler object.
- \_\_\_\_\_ 5. A calorie is the amount of energy required to raise the temperature of *one gram* of pure water by one degree Celsius.
- \_\_\_\_\_ 6. A *calorie* is the SI unit of heat and energy.
- \_\_\_\_\_ 7. The *specific heat* of a substance is the amount of heat required to raise the temperature of one gram of that substance by one degree Celsius.
- \_\_\_\_\_ 8. *Kinetic energy* is energy of motion.
- \_\_\_\_\_ 9. Chemicals participating in a chemical reaction contain only *potential energy*.
- \_\_\_\_\_ 10. One nutritional Calorie is equal to *100 calories*.
- \_\_\_\_\_ 11. One calorie equals *4.184 joules*.
- \_\_\_\_\_ 12. When a fuel is burned, some of its *chemical potential energy* is lost as heat.
- \_\_\_\_\_ 13. To convert kilojoules to joules, *divide* the number of kilojoules by 1000 joules/1 kilojoule.

Answer the following question. Show all your work.

14. If the temperature of a 500.0-g sample of liquid water is raised 2.00°C, how much heat is absorbed by the water? The specific heat of liquid water is 4.184 J/(g·°C).

# Chemistry Matter And Change Chapter 14 Study Guide

**Martha Mackin**



## **Chemistry Matter And Change Chapter 14 Study Guide:**

**Chemistry, Student Study Guide** James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH<sub>4</sub> from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

*Chemistry, Study Guide* James E. Brady, John R. Holum, 1993-02-11 Offers accurate lucid and interesting explanations of basic concepts and facts of chemistry while helping students develop skills in analytical thinking and problem solving Students are taught in a variety of ways to think of skills as tools that can be used to solve complex problems Several aids are included to help focus and inspire student interest frequent reference to common chemicals in commercial products numerous photographs of reactions in chapter practice exercises following worked examples

*Prentice Hall Chemistry*, 2000

**Student Study Guide to Accompany Petrucci's General Chemistry, 3rd. Ed** Robert K. Wismer, 1982

**Student Study Guide to Accompany Chemistry** Kenneth J. Hughes, Paul B. Kelter, 2002-05 Reviews essential points from the text presents problems and hints to assist the student in solving them Solutions to selected problems in the end of chapter problem sets are included By Kenneth Hughes of the University of Wisconsin Madison

**Study Guide for The Anatomy and Physiology Learning System** Edith Applegate, 2010-03-30 Designed to accompany The Anatomy and Physiology Learning System 4th Edition by Edith Applegate this study guide helps you learn and review basic A P concepts Each chapter emphasizes medical terminology with a set of key terms word parts clinical terms and abbreviations and then adds a variety of fun filled learning exercises review questions a quiz and a word puzzle The study guide corresponds to the textbook chapter for chapter Chapter learning objectives help you focus on the most important material Key concepts are defined on the first page of each chapter in the workbook Learning exercises for each chapter include short answer matching and diagrams to label and color Self quizzes allow you to measure your progress and understanding Fun and Games features end each chapter with a variety of engaging puzzles covering words and concepts A chapter summary provides a brief review of each chapter A chapter review provides questions for reinforcement and review of the concepts in each chapter

**CHEMISTRY SILBERBERG**, 2003

[Study Guide \[to Accompany\] General Chemistry](#) Dixie Goss, 2002

**Study Guide for E. Russell Hardwick's Introduction to Chemistry** Joan Bouillon, 1984

**Student Study Guide to Accompany Petrucci's General Chemistry** Robert K. Wismer, 1985

**Students Study Guide** Fredeen, Pearson, 2003-07

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**General Chemistry** Robert K. Wismer,1993

Student Study Guide and Solutions Manual to Accompany General, Organic, and Biochemistry Katherine J.

Denniston,Denniston Katherine,2006 *Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today* Seager/Slabaugh,2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application *Organic Chemistry* Francis A. Carey,2000 A Market Leading Traditional Approach to Organic Chemistry Throughout all seven editions Organic Chemistry has been designed to meet the needs of the mainstream two semester undergraduate organic chemistry course This best selling text gives students a solid understanding of organic chemistry by stressing how fundamental reaction mechanisms function and reactions occur With the addition of handwritten solutions new cutting edge molecular illustrations updated spectroscopy coverage seamless integration of molecular modeling exercises and state of the art multimedia tools the 7th edition of Organic Chemistry clearly offers the most up to date approach to the study of organic chemistry Study Guide [to] Chemistry, a Systematic Approach [by] Sisler, Dresdner, and Mooney William H. Myers,William T. Mooney,1980

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