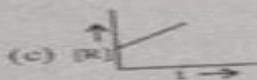
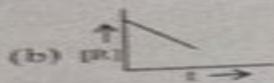


UNIT 3

CHEMICAL KINETICS

Multiple Choice Questions (MCQs)

- The reactions which occurs very fast are
 - mixing aqueous solutions of silver nitrate and sodium chloride
 - Rusting of iron.
 - Inversion of cane sugar
 - hydrolysis of starch.
- In the rate equation, when the conc. of reactants is unity then rate is equal to
 - specific rate constant
 - average rate constant
 - instantaneous rate constant
 - None of above
- The rate of reaction between two specific time intervals is called
 - instantaneous rate
 - average rate
 - specific rate
 - ordinary rate
- Which one of the following statements for the order of a reaction is incorrect ?
 - Order can be determined only experimentally.
 - Order is not influenced by stoichiometric coefficient of the reactants.
 - Order of reaction is sum of power to the concentration terms of reactants to express the rate of reaction.
 - Order of reaction is always whole number.
- Units of rate constant of first and zero order reactions in terms of molarity M unit are respectively
 - $\text{sec}^{-1}, \text{M sec}^{-1}$
 - $\text{sec}^{-1}, \text{M}$
 - $\text{M sec}^{-1}, \text{sec}^{-1}$
 - $\text{M}, \text{sec}^{-1}$
- In a 1st order reaction, reactant concentration C varies with time t as :
 - $1/C$ increases linearly with t
 - $\log C$ decreases linearly with t
 - C decreases with $1/t$
 - $\log C$ decreases with $1/t$
- For a first order reaction, a plot of $\log(a-x)$ against time is a straight line with a slope equal to
 - $\frac{-k}{2.303}$
 - $-2.303k$
 - $\frac{2.303}{k}$
 - $-\frac{E_a}{2.303R}$
- In a reaction, when the concentration of reactant is increased two times, the increase in rate of reaction was four times. Order of reaction is :
 - zero
 - 1
 - 2
 - 3
- The plot that represents the zero order reaction is :



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