

Formula Sheet For General Chemistry (Nov. 16 2007)  
 Blinn College Learning Center

DESCRIPTION	EQUATION
Ideal gas equation	$PV = nRT$
Adiabatic change	$PV = k$
Charles' Law	$\frac{V}{t} = k$
Bohr Radius	$a_0 = \frac{h^2}{m_e k e^2}$
Radii of stable orbits in the Bohr model	$r = n^2 \frac{h^2}{m_e k Z e^2} = n^2 \frac{a_0}{Z}$
Van der Waals equation	$\left( P + \frac{an^2}{V^2} \right) (V - bn) = nRT$
Entropy Change	$\Delta S^\circ = \sum S^\circ \text{products} - \sum S^\circ \text{reactants}$
Enthalpy Change	$\Delta H^\circ = \sum H_f^\circ \text{products} - \sum H_f^\circ \text{reactants}$
Gibb's Free Energy Change Defined	$\Delta G^\circ = \sum G_f^\circ \text{products} - \sum G_f^\circ \text{reactants}$
Gibb's Free Energy Change in Terms of Enthalpy, Absolute Temperature, and Entropy	$\Delta G^\circ = \Delta H^\circ - T \Delta S^\circ$
Gibb's Free Energy Change in Terms of Gas Constant, Absolute Temperature, and Equilibrium Constant	$\Delta G^\circ = -RT \ln K = -2.303RT \log K$
Gibb's Free Energy Change in Terms of Number of Moles, Faraday, and Standard Reduction Potential	$\Delta G^\circ = -n \mathfrak{F} E^\circ$
Reaction Quotient	$Q = \frac{[C]^c [D]^d}{[A]^a [B]^b}$ where $aA + bB \rightarrow cC + dD$
Electric Current	$I = \frac{q}{t}$
Cell Voltage	$E_{\text{cell}} = E_{\text{cell}}^\circ - \frac{RT}{n \mathfrak{F}} \ln Q = E_{\text{cell}}^\circ - \frac{0.0592}{n} \log Q$

# Chemical Engineering Formula Sheet

**Daniel F McAuley**



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