

Question number	Answer	Marks	Guidance															
1 (a)	C	B1																
1 (b)	B	B1																
1 (c)	D	B1																
2 (a)	38	B1																
2 (b)	32	B1																
2 (c)	5	B1																
3 (a) (i)	23	B1																
3 (a) (ii)	$^{56}_{26}\text{Fe}$	B2																
3 (b)	$^{23}_{11}\text{Na}$	B2																
3 (c)	$^{88}_{38}\text{Sr}$	B2																
4 (a)	Isotopes are atoms of the same element with different numbers of neutrons and different masses.	B1																
4 (b)	Relative isotopic mass is the mass of an isotope relative to one-twelfth of the mass of an atom of carbon-12.	B2																
4 (c)	71	B1																
4 (d)	<table border="1"> <thead> <tr> <th></th> <th>protons</th> <th>neutrons</th> <th>electrons</th> <th>charge</th> </tr> </thead> <tbody> <tr> <td>^{17}O</td> <td>8</td> <td>9</td> <td>10</td> <td>2-</td> </tr> <tr> <td>^{27}Al</td> <td>13</td> <td>14</td> <td>10</td> <td>3+</td> </tr> </tbody> </table>		protons	neutrons	electrons	charge	^{17}O	8	9	10	2-	^{27}Al	13	14	10	3+	B2	
	protons	neutrons	electrons	charge														
^{17}O	8	9	10	2-														
^{27}Al	13	14	10	3+														
5 (a)	Relative atomic mass is the weighted mean mass of an atom of an element relative to one-twelfth of the mass of an atom of carbon-12.	B3																
5 (b)	$A_r = 32.09$	B2																
5 (c)	<table border="1"> <thead> <tr> <th></th> <th>protons</th> <th>neutrons</th> <th>electrons</th> </tr> </thead> <tbody> <tr> <td>^{32}S</td> <td>16</td> <td>16</td> <td>16</td> </tr> <tr> <td>$^{32}\text{S}^{2-}$</td> <td>16</td> <td>16</td> <td>18</td> </tr> </tbody> </table>		protons	neutrons	electrons	^{32}S	16	16	16	$^{32}\text{S}^{2-}$	16	16	18	B1				
	protons	neutrons	electrons															
^{32}S	16	16	16															
$^{32}\text{S}^{2-}$	16	16	18															
6 (a)	$3\text{Mg(s)} + \text{N}_2\text{(g)} \rightarrow \text{Mg}_3\text{N}_2\text{(s)}$ 1 mark for correct species and state symbols 1 mark for balance	B1 × 2																
6 (b)	$\text{Ca(s)} + 2\text{H}_2\text{O(l)} \rightarrow \text{Ca(OH)}_2\text{(aq)} + \text{H}_2\text{(g)}$ 1 mark for correct species and state symbols 1 mark for balance	B1 × 2																
6 (c)	$6\text{NaOH(aq)} + \text{Fe}_2\text{(SO}_4\text{)}_3\text{(aq)} \rightarrow 2\text{Fe(OH)}_3\text{(s)} + 3\text{Na}_2\text{SO}_4\text{(aq)}$ 1 mark for correct species and state symbols 1 mark for balance	B1 × 2																

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Jin-Ying Zhang



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Table of Contents Skill Practice 27 Chemistry Answers

1. Understanding the eBook Skill Practice 27 Chemistry Answers
 - The Rise of Digital Reading Skill Practice 27 Chemistry Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Skill Practice 27 Chemistry Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Skill Practice 27 Chemistry Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Skill Practice 27 Chemistry Answers
 - Personalized Recommendations
 - Skill Practice 27 Chemistry Answers User Reviews and Ratings
 - Skill Practice 27 Chemistry Answers and Bestseller Lists
5. Accessing Skill Practice 27 Chemistry Answers Free and Paid eBooks

- Skill Practice 27 Chemistry Answers Public Domain eBooks
 - Skill Practice 27 Chemistry Answers eBook Subscription Services
 - Skill Practice 27 Chemistry Answers Budget-Friendly Options
6. Navigating Skill Practice 27 Chemistry Answers eBook Formats
- ePub, PDF, MOBI, and More
 - Skill Practice 27 Chemistry Answers Compatibility with Devices
 - Skill Practice 27 Chemistry Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Skill Practice 27 Chemistry Answers
 - Highlighting and Note-Taking Skill Practice 27 Chemistry Answers
 - Interactive Elements Skill Practice 27 Chemistry Answers
8. Staying Engaged with Skill Practice 27 Chemistry Answers
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Skill Practice 27 Chemistry Answers
9. Balancing eBooks and Physical Books Skill Practice 27 Chemistry Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Skill Practice 27 Chemistry Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Skill Practice 27 Chemistry Answers
- Setting Reading Goals Skill Practice 27 Chemistry Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Skill Practice 27 Chemistry Answers
- Fact-Checking eBook Content of Skill Practice 27 Chemistry Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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