

Lab 22A

Investigating Chemical Equilibrium

Purpose: To investigate the macroscopic properties of two chemical systems at equilibrium, observe shifts in equilibrium concentrations as a function of concentration and temperature and relate the to systems, which are exploring the structures through applying Le Chatelier's principle.

Proble:

1. Write a balanced equation for the reaction in part 1.
$$\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$$
2. Write a balanced equation for the reaction in part 2.
$$\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$$
3. Write a balanced equation for the reaction in part 3.
$$\text{FeSCN}^{2+}(\text{aq}) + \text{H}^{+}(\text{aq}) \rightleftharpoons \text{FeSCNH}^{3+}(\text{aq})$$
4. In part 4, what is the purpose of adding potassium thiocyanate to 100 mL of 0.1 M FeCl₃ solution?
To observe shifts in the equilibrium and to see the effect of adding a reactant. The purpose of adding FeCl₃ is to observe the effect of adding a reactant and to observe the effect of adding a product.
5. What is the purpose of adding distilled water in part 5?
To observe the effect of adding a reactant and to observe the effect of adding a product.
6. What is the purpose of adding HCl in part 6?
To observe the effect of adding a reactant and to observe the effect of adding a product.
7. Write the equation for the reaction between Fe³⁺ and OH⁻ ions.
$$\text{Fe}^{3+}(\text{aq}) + \text{OH}^{-}(\text{aq}) \rightleftharpoons \text{Fe}(\text{OH})_3(\text{s})$$
8. Write the equation for the reaction between Fe³⁺ and OH⁻ ions.
$$\text{Fe}^{3+}(\text{aq}) + \text{OH}^{-}(\text{aq}) \rightleftharpoons \text{Fe}(\text{OH})_3(\text{s})$$
9. What is the purpose of adding HCl in part 9?
To observe the effect of adding a reactant and to observe the effect of adding a product.

Investigating Chemical Equilibrium Lab 12a Answer

SB Merriam



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