

## Covalent Bonding Review Sheet – Dec 12, 2012

Name \_\_\_\_\_ class period \_\_\_\_\_

- Given the formula, a compound can be named using conventional systems that include Greek prefixes and Roman numeral where appropriate.
- Differences in electronegativity values can be used to predict where a bond fits on the continuum between ionic and covalent bonds.
- Since most compounds contain multiple bonds, a substance may contain more than one type of bond.

### Section 1: Naming and Formula Writing

Write the formulas for the following covalent compounds

1. Nitrogen Monoxide NO
2. Dinitrogen Monoxide N<sub>2</sub>O
3. Nitrogen Dioxide NO<sub>2</sub>
4. Dinitrogen Trioxide N<sub>2</sub>O<sub>3</sub>
5. Dinitrogen Tetraoxide N<sub>2</sub>O<sub>4</sub>
6. Dinitrogen Pentaoxide N<sub>2</sub>O<sub>5</sub>
7. carbon dioxide CO<sub>2</sub>
8. phosphorus triiodide PI<sub>3</sub>
9. sulfur dichloride SCl<sub>2</sub>
10. nitrogen trifluoride NF<sub>3</sub>

Write names from the following formulas

11. N<sub>2</sub>F<sub>4</sub> dinitrogen tetrafluoride
12. SCl<sub>4</sub> sulfur tetrachloride
13. ClF<sub>3</sub> chlorine trifluoride
14. SiO<sub>2</sub> silicon dioxide
15. P<sub>4</sub>O<sub>10</sub> tetraphosphorus decaoxide
16. SiF<sub>4</sub> silicon tetrafluoride
17. H<sub>2</sub>O dihydrogen monoxide
18. Br<sub>2</sub> dibromide or bromine
19. HBr hydrogen monobromide
20. N<sub>2</sub>S<sub>3</sub> dinitrogen trisulfide

# Concept Review Section Covalent Bonds Answer Key

**JG Myers**



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