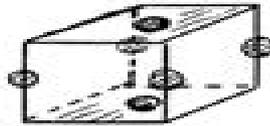


PROBLEM SOLUTIONS
CHAPTER 1

(1.1)
(a)



... $\frac{1}{2}$ atom inside cell on top and bottom faces
... $\frac{1}{4}$ atom along each vertical edge

$$\text{TOTAL ATOMS} = 2 \times \frac{1}{2} + 4 \times \frac{1}{4} = \underline{\underline{2}}$$

(b)



⇐ TOP VIEW

- ... plane #1
- △ ... plane #2
- ... plane #3

(c) bcc

(1.2)

(a) The diamond lattice is composed of two interpenetrating fcc lattices. Each fcc lattice unit cell contains 4 atoms. ($\frac{1}{2}$ atom at each of 6 faces and $\frac{1}{8}$ atom at each of 8 corners.) Thus the Si lattice contains 8 ATOMS per unit cell. Inspecting Fig. 1.5a leads to the same conclusion -- there are 4 atoms totally inside the cube and an fcc arrangement on the faces of the cube.

(b)

$$\frac{\text{Si ATOMS}}{\text{CM}^3} = \frac{\text{ATOMS/UNIT CELL}}{a^3} = \frac{8}{(5.431 \times 10^{-8})^3} = 4.997 \times 10^{22} \checkmark$$

↙ from Table 1.5

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Table of Contents Advanced Semiconductor Fundamentals Instructor Manual

1. Understanding the eBook Advanced Semiconductor Fundamentals Instructor Manual
 - The Rise of Digital Reading Advanced Semiconductor Fundamentals Instructor Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Advanced Semiconductor Fundamentals Instructor Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Advanced Semiconductor Fundamentals Instructor Manual
 - User-Friendly Interface
4. Exploring eBook Recommendations from Advanced Semiconductor Fundamentals Instructor Manual
 - Personalized Recommendations
 - Advanced Semiconductor Fundamentals Instructor Manual User Reviews and Ratings
 - Advanced Semiconductor Fundamentals Instructor Manual and Bestseller Lists

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 - Advanced Semiconductor Fundamentals Instructor Manual Public Domain eBooks
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 - ePub, PDF, MOBI, and More
 - Advanced Semiconductor Fundamentals Instructor Manual Compatibility with Devices
 - Advanced Semiconductor Fundamentals Instructor Manual Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Advanced Semiconductor Fundamentals Instructor Manual
 - Highlighting and Note-Taking Advanced Semiconductor Fundamentals Instructor Manual
 - Interactive Elements Advanced Semiconductor Fundamentals Instructor Manual
8. Staying Engaged with Advanced Semiconductor Fundamentals Instructor Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Advanced Semiconductor Fundamentals Instructor Manual
9. Balancing eBooks and Physical Books Advanced Semiconductor Fundamentals Instructor Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Advanced Semiconductor Fundamentals Instructor Manual
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Advanced Semiconductor Fundamentals Instructor Manual
 - Setting Reading Goals Advanced Semiconductor Fundamentals Instructor Manual
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Advanced Semiconductor Fundamentals Instructor Manual
 - Fact-Checking eBook Content of Advanced Semiconductor Fundamentals Instructor Manual
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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