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# Mathematical Statistics

with Applications

7th Edition



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**Puja Mehta**



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web mar 14 2023 the ultimate goal for balancing chemical equations is to make both sides of the reaction the reactants and the products equal in the number of atoms per element this stems from the universal law of the conservation of mass which states that matter can neither be created nor destroyed

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web balancing equations the chemical equation described in section 4 1 is balanced meaning that equal numbers of atoms for each element involved in the reaction are represented on the reactant and product sides this is a requirement the equation must satisfy to be consistent with the law of conservation of matter

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web because of the law of conservation of mass balancing chemical equations gives a more accurate representation on what is happening in a reaction the accuracy of the ratios of moles is important in calculating the theoretical yield and how much stuff is reacting mass wise stoichiometry

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web a key step in balancing chemical equations is correctly identifying the formulas of the reactants and products for example consider the reaction between calcium oxide  $\text{CaO}$  and  $\text{H}_2\text{O}$  to form aqueous calcium hydroxide  $\text{Ca(OH)}_2$  is it possible to balance the equation if you incorrectly identify the product as  $\text{CaOH}$  and if so what is the equation

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web balance the following chemical equation  $\text{Mg(OH)}_2 + \text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2\text{O}$  note all reactants and products require a coefficient of at least one stuck review related articles videos or use a hint report a problem do 4 problems

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