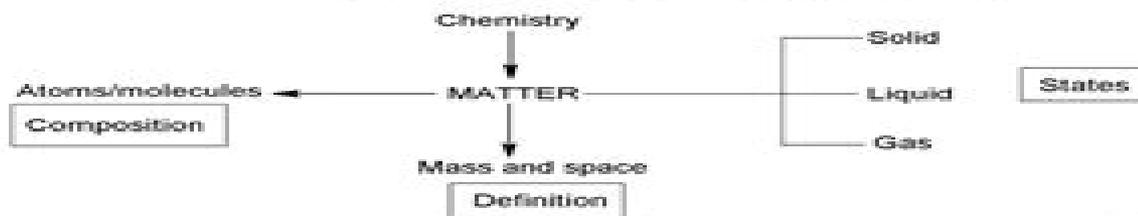


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**Atoms:** Building block of matter. It may or may not exist free in nature

Examples: H, He, Na, C



**Nucleus:** It is central part of an atom.

**Shells:** Circular paths around the nucleus where electrons reside. They are also called orbits or energy levels.

No. of es in a shell =  $2n^2$

Where n = no of shell.

1<sup>st</sup> shell = 2es, 2<sup>nd</sup> shell = 8es, 3<sup>rd</sup> shell = 18es.

At least 8 electron should be present in last for the stability of an atom. However if an atom has one shell like Hydrogen and Helium then 2 electrons are required for complete shell and stability.

**Atomic number:**

Number of protons present in the nucleus of an atom of an element.

Examples:  $_{11}\text{Na}$ ,  $_6\text{C}$ ,  $_8\text{O}$

**Atomic mass or Mass number or nucleon number:** It is the sum of protons and neutrons present in the nucleus of an atom of an element.

Examples:  $^{23}\text{Na}$ ,  $^{12}\text{C}$ ,  $^{16}\text{O}$

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**R Barnett**



## **O Level Chemistry Notes:**

## **O Level Chemistry Notes** Book Review: Unveiling the Power of Words

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