

Lab – Shapes of Covalent Molecules

Introduction

The type of chemical bond that will form between two atoms can be predicted by calculating the difference in the atoms' **electronegativities**. When the values of two atoms' electronegativities are far apart, one atom loses one or more electrons to the other and an **ionic bond** is formed. When two atoms' electronegativities are closer in value, the atoms share pairs of electrons and a **covalent bond** is formed. Covalent bonds are considered **nonpolar** when the atoms' electronegativities are nearly the same and the electrons are shared evenly between the atoms. Covalent bonds are considered **polar** when the two atoms unequally share electron pairs.

Electron pairs around atoms, whether they are bonding or nonbonding (lone) pairs, all repel each other and try to move as far from each other as possible. According to **VSEPR theory** (valence shell electron pair repulsion theory), this arrangement of valence electron pairs determines the shape of covalent molecules.

When a covalent molecule contains only nonpolar bonds, the molecule as a whole is also nonpolar. Its electron cloud is evenly and symmetrically spread around the molecule so that there are no areas of partially positive or negative charge. When a covalent molecule contains polar bonds and the electron cloud is not symmetric, the molecule as a whole is polar and has areas of partial charge (dipoles). The molecule is partially negative (δ^-) near the more electronegative atom and partially positive (δ^+) near the less electronegative atom. In some cases, covalent molecules containing polar bonds are, as a whole, nonpolar if the "pulls" are symmetric.

Procedure

1. Construct the electron dot (Lewis) structure for the molecule given.
2. Determine the shape of the molecule by constructing the molecule.
3. Record the electronegativities of the elements and calculate the difference.
4. State the type of bond (polar covalent or nonpolar covalent) formed using the data provided by the teacher.
5. State whether the molecule as a whole is polar or nonpolar by asking yourself, "If there were a tug of war in this molecule, would anyone win?" If there is a winner, the molecule is polar, if there would be a tie, the molecule is nonpolar.

Formula	Electron dot structure	Shape of molecule	Calculate electronegativity difference	Bond (polar or nonpolar)	Molecule (polar or nonpolar)
H_2					
F_2					

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