



E1 reading plz

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51 ~ 60	Recommended Reading 1	3 ~ 4
61 ~ 70	Underlining 1	3 ~ 4
71 ~ 80	Underlining 2	3 ~ 4
81 ~ 90	Underlining 3	3 ~ 4
91 ~ 100	Underlining 4	3 ~ 4
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151 ~ 160	Reason and Result 5	3 ~ 4
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171 ~ 180	Vocabulary Review 1	3 ~ 4
181 ~ 190	Vocabulary Review 2	3 ~ 4
191 ~ 200	Recommended Reading 3	3 ~ 4

\*SCT = Standard Completion Time

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EII 1-2-3-4

<b>1</b>	<b>2</b>
<p><b>I</b></p> <ol style="list-style-type: none"> <li>his mouth peach juice his tongue</li> <li>every few seconds bite out of the wall</li> <li>solid wall stone in the middle of the peach</li> </ol>	<p><b>I</b></p> <ol style="list-style-type: none"> <li>5</li> <li>2</li> <li>1</li> <li>3</li> <li>4</li> </ol> <p><b>II</b></p> <ol style="list-style-type: none"> <li>went to work in Africa</li> <li>joined the British Royal Air Force as a fighter pilot</li> <li>(James and the Giant Peach) (Mr. Dahl's first children's book)</li> </ol>
<b>3</b>	<b>4</b>
<p><b>I</b></p>	<p><b>I</b></p>

# Kumon Answer Level Cii

**M Tight**



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web limiting reactant theoretical and percentage yields key  $4 KO_2 + 2 H_2O \rightarrow 4 KOH + 3 O_2$  g if a reaction vessel contains 0 mol  $KO_2$  and 0 mol  $H_2O$  what is the limiting reactant

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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu}(\text{NO}_3)_2 + 2\text{Ag}$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8 3 mole  $\text{H}_2\text{S}$  and multiply it with 2 mole  $\text{H}_2\text{O}$  over 2 mole  $\text{H}_2\text{S}$  from the equation  

$$8 3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} = 8 3 \text{ mole } \text{H}_2\text{O}$$
  
 8 3 mole  $\text{H}_2\text{S}$  x 18g  $\text{H}_2\text{O}$  / 1 mole  $\text{H}_2\text{O}$  149 4g  $\text{H}_2\text{O}$  is your theoretical yield so percent yield 137 1g  $\text{H}_2\text{O}$  actual yield 149 4g  $\text{H}_2\text{O}$  x 100  
 91 77 problems 6 8 are similar to problem 5

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web 1 csun edu hcchm001 introchemhandouts html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of  $\text{NH}_3$  can be produced theoretically from the reaction of 5 0 g of  $\text{N}_2$  what is the limiting reagent if 8 52 g are actually formed what is the

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$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$
  
 if the reaction of 21 3 grams of  $\text{O}_2$  produces a 47 3 yield how many grams of  $\text{CO}$  would be produced step one figure out how many moles of the limiting reagent you have  

$$\frac{21 3 \text{ grams } \text{O}_2}{32 \text{ g/mole}} = 0 66 \text{ moles } \text{O}_2$$

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web a write a balanced equation for the reaction  

$$\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$$
  
 4 al 3 o2 2 al2o3 b determine the theoretical yield of  $\text{Al}_2\text{O}_3$  160 0

g al x 1 mol al x 2 mol al<sub>2</sub>O<sub>3</sub> x 101 964 g al<sub>2</sub>O<sub>3</sub> 302 3 g al<sub>2</sub>O<sub>3</sub> 26 982 g al 4 mol al 1 mol al<sub>2</sub>O<sub>3</sub> c determine the percent yield  
 yield actual x 100 260 0 g x 100 86 01 theoretical 302 3 g

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web 7 na 2 c 2 o 4 is the limiting reactant percent yield 86 6 8 only four molecules can be made 9 this amount cannot be  
 weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12  
 c is the limiting reagent 4 33 g of h 2 are left over 13

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 from a set of reactants calculate how much product will be produced from the limiting reagent calculate how much reactant s  
 remains when the reaction is complete

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web the percent yield is calculated as follows text percent yield  $\frac{\text{text actual yield}}{\text{text theoretical yield}} \times 100$  percent  
 yield theoretical yield actual yield 100 based on this definition we would expect a percent yield

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 how much product will be produced from the limiting reactant calculate how much reactant s remains when the reaction is  
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 ch<sub>3</sub>oh can be produced by the following reaction 2h<sub>2</sub> co ch<sub>3</sub>oh calculate the theoretical yield of ch<sub>3</sub>oh if 68 5 g of co is  
 reacted with 8 6 g of h<sub>2</sub>

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 which reactant is limiting assuming we started with 30 0 grams of ammonium nitrate and 50 0 grams of sodium phosphate  
 what is the

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80 g  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2$  x 1 333 8 g  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2\text{O}_5$  28 g  $\text{CO}$  1 mol  $\text{CO}$

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web Jun 2 2020 the percent yield of a reaction is the ratio of the actual yield to the theoretical yield multiplied by 100 to give a percentage  
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