

## Flame Test Lab

### Activity Key

#### Background

The heat from a laboratory burner will cause the ions of some elements to give off light. Electrons will absorb the heat energy from the flame and will “jump” to a higher energy level. When the electrons return to their original energy levels, this absorbed energy is released as light. Different elements absorb and release different amounts of energy and thus produce different kinds of light. This light may be visible or invisible, infrared or ultraviolet radiation. Analysis of this light can be used to identify elements and molecules. In today’s lab, we will observe the visible range of light emission.

#### Objectives

1. Perform a flame test to identify the characteristic color of metal ions.
2. Identify the metal ion in an unknown solution.
3. Calculate the energy of emitted photons.

#### Materials: (per lab group)

10 ml dilute solutions of the following:

Barium Chloride  
Calcium Chloride  
Lithium Chloride  
Potassium Chloride  
Sodium Chloride  
Strontium Chloride  
Unknown Solution

7 Wood Splints  
1 Bunsen Burner  
Safety Goggles & Aprons

#### Note:

If chloride compounds are not available, metal nitrate compounds may be substituted.

Use dilute or approximately 0.1 M solutions

#### Unknowns:

Number the beakers 1-12.  
Add solutions as follows:

- |                       |                        |
|-----------------------|------------------------|
| 1. Strontium Chloride | 7. Calcium Chloride    |
| 2. Calcium Chloride   | 8. Strontium Chloride  |
| 3. Potassium Chloride | 9. Barium Chloride     |
| 4. Barium Chloride    | 10. Potassium Chloride |
| 5. Lithium Chloride   | 11. Sodium Chloride    |
| 6. Sodium Chloride    | 12. Lithium Chloride   |

#### Procedure

1. Put on your safety goggles and an apron.
2. Obtain a wood splint and the barium chloride solution.
3. Dip the wood splint into the solution and allow it to soak for 1-2 minutes.
4. Light Bunsen Burner. USE CAUTION!
5. Remove the splint from the solution and hold the end of the splint into the flame. Do not burn the splint.
6. Observe and record the color of the flame.
7. Repeat steps 2-6 using the remaining chloride solutions.
8. Obtain an unknown solution from your teacher. Record the number of your unknown.
9. Soak a wood splint in the unknown solution.
10. Hold the splint in the flame and record the color of the flame that is produced.
11. Using your data, identify the metal ion in your unknown solution.

# Chemistry Flame Test Lab Answers

**Ensheng Dong**



## **Chemistry Flame Test Lab Answers:**

*Laboratory Manual for Principles of General Chemistry* Jo Allan Beran, 2010-11-01 This new edition of the Beran lab manual emphasizes chemical principles as well as techniques The manual helps students understand the timing and situations for the various techniques The Beran lab manual has long been a market leading lab manual for general chemistry Each experiment is presented with concise objectives a comprehensive list of techniques and detailed lab intros and step by step procedures

*Chemistry in the Laboratory* James M. Postma, Julian L. Robert, J. Leland Hollenberg, 2004-03-12 This clearly written class tested manual has long given students hands on experience covering all the essential topics in general chemistry Stand alone experiments provide all the background introduction necessary to work with any general chemistry text This revised edition offers new experiments and expanded information on applications to real world situations

**Lab Experiments for Modern Chemistry** Tzimopoulo, 1990

**Holt Chemistry** Ralph Thomas Myers, 2004

*Laboratory Exercises to Accompany First Principles of Chemistry* Raymond Bedell Brownlee, 1908

**Big Book of Home Learning** Mary Pride, 1991-07 Learn at home with exciting products for all school subjects New

**Chemistry** Karen Timberlake, 1999 Suitable for one or two term lab courses covering general organic and biological chemistry this new edition written by Karen Timberlake features many improvements to the insightful experiments that have made it the leading lab manual Each experiment encourages critical thinking with laboratory goals discussion of related concepts clear instructions new pre lab questions and comprehensive report pages Forty one experiments illustrate the basic principles of chemistry

General Chemistry Ralph H. Petrucci, 2007 General Chemistry Principles and Modern Applications is recognized for its superior problems lucid writing precision of argument and precise and detailed treatment of the subject Popular and innovative features include Feature Problems follow up A and B Practice Exercises to accompany every in chapter Example Focus On application boxes and new Keep in Mind marginal notes Every new copy of the Ninth Edition comes with a Student MediaPak which includes access to the Companion Website with GradeTracker available at <http://www.prenhall.com/petrucci> the Student Accelerator CD and the Virtual ChemLab Workbook and CD This package includes Basic Media Pack Wrap Companion WEBSITE Grade Tracker Access Code Card Virtual ChemLab General Chemistry Student Lab Manual Workbook

**Exercises in Chemistry, Systematically Arranged to Accompany "An Elementary Study of Chemistry", Introductory College Course, by William Pherson and William Edwards Henderson** William McPherson, 1925

General, Organic, and Biological Chemistry Study Guide and Selected Solutions Karen C. Timberlake, 2001-11 Keyed to the learning goals in the text this guide is designed to promote active learning through a variety of exercises with answers and mastery exams The guide also contains complete solutions to odd numbered problems

*A Laboratory Text Book of Practical Chemistry* William George Valentin, 1871

Experiments in General Chemistry Kenneth W. Whitten, Carl B. Bishop, Kenneth D. Gailey, 1992

**Chemistry** John S. Phillips, Cheryl Wistrom, 2000

**Nuclear Science Abstracts**, 1958

The Journal of Industrial and Engineering Chemistry ,1913      *Journal of Industrial and Engineering Chemistry* ,1922  
*Instructors Manual to Lab Manual* Ralph Petrucci,William Harwood,Geoffrey Herring,2001      Gradwohl Laboratory  
Digest ,1964-07      *Virtual Chemlab* Brian F. Woodfield,Matthew C. Asplund,Steven Haderlie,2006 Contains a full virtual lab  
environment as well as the pre arranged labs that are refer enced in the workbook and at the end of the chapter in the  
textbook Virtual ChemLab can be run directly from the CD or installed on the student s computer      Directory of Testing  
Laboratories ,1992

## **Chemistry Flame Test Lab Answers** Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Chemistry Flame Test Lab Answers**," published by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we will delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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