

## NUCLEAR CHEMISTRY

### Practice Problems

In your notebook, solve the following problems.

#### SECTION 25.1 NUCLEAR RADIATION

- What happens to the mass number and atomic number of an atom that undergoes beta decay?
- A nuclide of an element undergoes alpha particle decay. How do the atomic number and mass number of the particle change?
- Give the composition of the nucleus of the following isotopes.
  - ${}^{234}\text{Th}$
  - ${}^{232}\text{Th}$
  - ${}^{210}\text{At}$
- Complete each of the following equations.
  - ${}^{234}\text{Th} \rightarrow \text{?} + \text{?}$
  - ${}^{210}\text{At} \rightarrow \text{?} + \text{?}$
  - ${}^{234}\text{Th} \rightarrow \text{?} + \text{?}$

#### SECTION 25.2 NUCLEAR TRANSFORMATIONS

- Write a nuclear equation for the following radioactive processes.
  - alpha decay of francium-223
  - electron capture by beryllium-7
  - beta emission by argon-37
  - positron emission by boron-12
- Complete the equations for these transmutation reactions.
  - ${}^{24}\text{Mg} + \text{?} \rightarrow \text{?} + \text{?}$
  - ${}^{238}\text{U} + \text{?} \rightarrow \text{?} + \text{?} + \text{?}$
  - ${}^{24}\text{Mg} + \text{?} \rightarrow \text{?} + \text{?} + \text{?}$
  - ${}^{238}\text{U} \rightarrow \text{?} + \text{?} + \text{?} + \text{?}$
  - $\text{?} + \text{?} \rightarrow \text{?} + \text{?} + \text{?} + \text{?}$
- Polonium-214 has a relatively short half-life of 264 s. How many seconds would it take for 6.0 g of this isotope to decay to 0.25 g?
- How many days does it take for 16 g of gadolinium-153 to decay to 1.0 g? The half-life of gadolinium-153 is 17 days.
- By approximately what factor would the mass of a sample of copper-66 decrease in 10 minutes? The half-life of copper-66 is 5.49 min.
- In 5.49 seconds, 1.00 g of argon-39 decay to leave only 0.10 g. What is the half-life of argon-39?

# Chapter 25 Nuclear Chemistry Pearson Answer Key

**Scott C. Dulebohn**



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## **Table of Contents Chapter 25 Nuclear Chemistry Pearson Answer Key**

1. Understanding the eBook Chapter 25 Nuclear Chemistry Pearson Answer Key
  - The Rise of Digital Reading Chapter 25 Nuclear Chemistry Pearson Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 25 Nuclear Chemistry Pearson Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 25 Nuclear Chemistry Pearson Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 25 Nuclear Chemistry Pearson Answer Key

- Personalized Recommendations
  - Chapter 25 Nuclear Chemistry Pearson Answer Key User Reviews and Ratings
  - Chapter 25 Nuclear Chemistry Pearson Answer Key and Bestseller Lists
5. Accessing Chapter 25 Nuclear Chemistry Pearson Answer Key Free and Paid eBooks
    - Chapter 25 Nuclear Chemistry Pearson Answer Key Public Domain eBooks
    - Chapter 25 Nuclear Chemistry Pearson Answer Key eBook Subscription Services
    - Chapter 25 Nuclear Chemistry Pearson Answer Key Budget-Friendly Options
  6. Navigating Chapter 25 Nuclear Chemistry Pearson Answer Key eBook Formats
    - ePub, PDF, MOBI, and More
    - Chapter 25 Nuclear Chemistry Pearson Answer Key Compatibility with Devices
    - Chapter 25 Nuclear Chemistry Pearson Answer Key Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Chapter 25 Nuclear Chemistry Pearson Answer Key
    - Highlighting and Note-Taking Chapter 25 Nuclear Chemistry Pearson Answer Key
    - Interactive Elements Chapter 25 Nuclear Chemistry Pearson Answer Key
  8. Staying Engaged with Chapter 25 Nuclear Chemistry Pearson Answer Key
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Chapter 25 Nuclear Chemistry Pearson Answer Key
  9. Balancing eBooks and Physical Books Chapter 25 Nuclear Chemistry Pearson Answer Key
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Chapter 25 Nuclear Chemistry Pearson Answer Key
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Chapter 25 Nuclear Chemistry Pearson Answer Key
    - Setting Reading Goals Chapter 25 Nuclear Chemistry Pearson Answer Key
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Chapter 25 Nuclear Chemistry Pearson Answer Key

- Fact-Checking eBook Content of Chapter 25 Nuclear Chemistry Pearson Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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