

Algebra 1 Chapter 11 Practice Test

Replace the ___?___ with <, >, or = to make a true statement.

1. $\frac{17}{23}$ _____ $\frac{15}{19}$

2. $\frac{197}{6}$ _____ $33\frac{2}{7}$

Arrange each group of numbers in order from least to greatest.

3. $-\frac{39}{8}$, -4.7 , $-\frac{41}{9}$

4. $\frac{5}{7}$, $2\frac{11}{3}$, $\frac{12}{15}$, $\frac{12}{17}$

Find the number halfway between the given numbers.

5. $\frac{27}{41}$ and $\frac{31}{37}$

6. $-5\frac{2}{7}$ and $9\frac{1}{4}$

Express each rational number as a terminating or repeating decimal.

7. $\frac{4}{9}$

8. $3\frac{11}{20}$

9. $\frac{41}{55}$

Express each rational number as a fraction in simplest form.

10. 0.77

11. $0.\overline{6}$

12. $2.\overline{37}$

Express both numbers as fractions. Then find their product.

13. $\frac{2}{5}$ and 0.63

14. $0.\overline{4}$ and $\frac{2}{3}$

Algebra 1 Chapter 11

Zhe Xu



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proceed with the smog test do not continue with the test

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web video answer so now we'll talk about problems 58 from chapter four this problem whereas how do isotopes differ and how are they similar isotopes for the same element so isotopes differ in the number of neutrons

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web aug 19 2022 atoms with the same number of protons but different numbers of neutrons are called isotopes they share almost the same chemical properties but differ in mass and therefore in physical properties there are stable isotopes which do not emit radiation and there are unstable isotopes which do emit radiation

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web jan 5 2016 explanation isotopes are atoms of the same element same atomic number i e same number of protons in the nucleus but different atomic masses different mass numbers due to different numbers of neutrons in the nucleus for example we may get an isotope of chlorine with 17 protons and 18 neutrons hence atomic mass 35

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web may 16 2018 explanation an atom is characterized as a particular element by the number of positively charged massive nuclear particles i e z 1 we got hydrogen z 2 we got helium z 3 lithium z 47 we got silver the nucleus can also contain neutrons massive nuclear particles of zero charge

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web isotopes are distinct nuclear species or nuclides as technical term of the same chemical element they have the same atomic number number of protons in their nuclei and position in the periodic table and hence belong to the same chemical element but differ in nucleon numbers mass numbers due to different numbers of neutrons in

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web verified by toppr correct option is d atoms having same atomic number but different atomic mass are known as isotopes the difference in the mass is due to the difference in their number of neutrons with the same number of electrons and protons

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