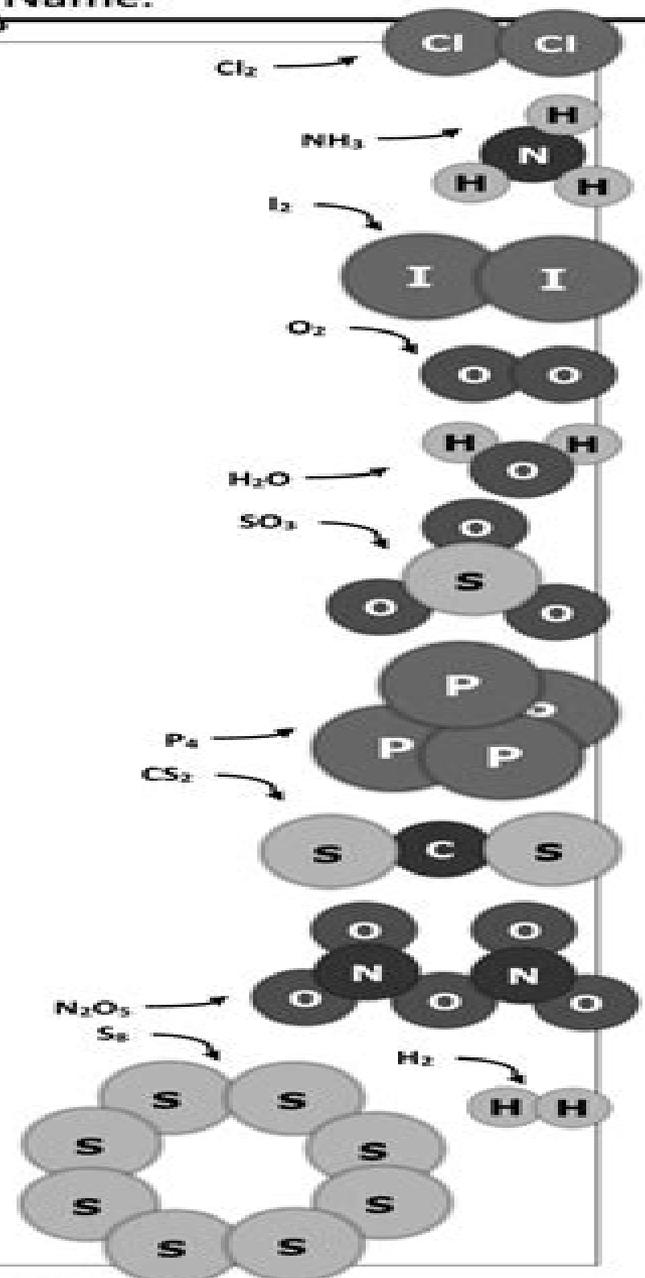


Balance the following chemical equations.

1.  $\underline{\quad}$  Na +  $\underline{\quad}$  Cl<sub>2</sub> →  $\underline{\quad}$  NaCl
2.  $\underline{\quad}$  Ca +  $\underline{\quad}$  O<sub>2</sub> →  $\underline{\quad}$  CaO
3.  $\underline{\quad}$  Na +  $\underline{\quad}$  O<sub>2</sub> →  $\underline{\quad}$  Na<sub>2</sub>O
4.  $\underline{\quad}$  N<sub>2</sub> +  $\underline{\quad}$  H<sub>2</sub> →  $\underline{\quad}$  NH<sub>3</sub>
5.  $\underline{\quad}$  K +  $\underline{\quad}$  I<sub>2</sub> →  $\underline{\quad}$  KI
6.  $\underline{\quad}$  Al +  $\underline{\quad}$  O<sub>2</sub> →  $\underline{\quad}$  Al<sub>2</sub>O<sub>3</sub>
7.  $\underline{\quad}$  H<sub>2</sub> +  $\underline{\quad}$  O<sub>2</sub> →  $\underline{\quad}$  H<sub>2</sub>O
8.  $\underline{\quad}$  H<sub>2</sub>O<sub>2</sub> →  $\underline{\quad}$  H<sub>2</sub>O +  $\underline{\quad}$  O<sub>2</sub>
9.  $\underline{\quad}$  Li +  $\underline{\quad}$  Cl<sub>2</sub> →  $\underline{\quad}$  LiCl
10.  $\underline{\quad}$  SO<sub>2</sub> +  $\underline{\quad}$  O<sub>2</sub> →  $\underline{\quad}$  SO<sub>3</sub>
11.  $\underline{\quad}$  S<sub>8</sub> +  $\underline{\quad}$  O<sub>2</sub> →  $\underline{\quad}$  SO<sub>3</sub>
12.  $\underline{\quad}$  P +  $\underline{\quad}$  O<sub>2</sub> →  $\underline{\quad}$  P<sub>2</sub>O<sub>5</sub>
13.  $\underline{\quad}$  P<sub>4</sub> +  $\underline{\quad}$  O<sub>2</sub> →  $\underline{\quad}$  P<sub>2</sub>O<sub>5</sub>
14.  $\underline{\quad}$  N<sub>2</sub> +  $\underline{\quad}$  F<sub>2</sub> →  $\underline{\quad}$  NF<sub>3</sub>
15.  $\underline{\quad}$  S<sub>8</sub> +  $\underline{\quad}$  C →  $\underline{\quad}$  CS<sub>2</sub>
16.  $\underline{\quad}$  N<sub>2</sub> +  $\underline{\quad}$  O<sub>2</sub> →  $\underline{\quad}$  N<sub>2</sub>O<sub>5</sub>
17.  $\underline{\quad}$  Mg +  $\underline{\quad}$  Cl<sub>2</sub> →  $\underline{\quad}$  MgCl<sub>2</sub>
18.  $\underline{\quad}$  Rb +  $\underline{\quad}$  P →  $\underline{\quad}$  Rb<sub>3</sub>P
19.  $\underline{\quad}$  Fe +  $\underline{\quad}$  O<sub>2</sub> →  $\underline{\quad}$  Fe<sub>2</sub>O<sub>3</sub>
20.  $\underline{\quad}$  K +  $\underline{\quad}$  Br<sub>2</sub> →  $\underline{\quad}$  KBr
21.  $\underline{\quad}$  HgO →  $\underline{\quad}$  O<sub>2</sub> +  $\underline{\quad}$  Hg
22.  $\underline{\quad}$  H<sub>2</sub> +  $\underline{\quad}$  Cl<sub>2</sub> →  $\underline{\quad}$  HCl
23.  $\underline{\quad}$  Zn +  $\underline{\quad}$  S<sub>8</sub> →  $\underline{\quad}$  ZnS
24.  $\underline{\quad}$  Ag<sub>2</sub>O →  $\underline{\quad}$  O<sub>2</sub> +  $\underline{\quad}$  Ag



# Pearson Education Chemical Reactions Packet Answers

**Emilie Sanchez**



**Pearson Education Chemical Reactions Packet Answers:**

## The Enigmatic Realm of **Pearson Education Chemical Reactions Packet Answers**: Unleashing the Language is Inner Magic

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