

## 16

## SOLUTIONS

**SECTION 16.1 PROPERTIES OF SOLUTIONS (pages 471–477)**

*This section identifies the factors that affect the solubility of a substance and determine the rate at which a solute dissolves.*

**► Solution Formation (pages 471–472)**

Look at Figure 16.1 on page 471 to help you answer Questions 1 and 2.

- Underline the condition that causes sugar to dissolve *faster* in water.
  - as a whole cube                      or            in granulated form?
  - when allowed to stand            or            when stirred?
  - at a higher temperature            or            a lower temperature?
- Name three factors that influence the rate at which a solute dissolves in a solvent.
  - agitation (stirring or shaking)
  - temperature
  - particle size
- Is the following sentence true or false? Finely ground particles dissolve more rapidly than larger particles because finer particles expose a greater surface area to the colliding solvent molecules.            true

**► Solubility (pages 472–473)**

- Complete the following table showing the steps in a procedure to determine the total amount of sodium chloride that will dissolve in 100 g of water at 25°C.

Procedure	Amount Dissolved	Amount Not Dissolved
Add 36.0 g of sodium chloride to the water	36.0 g	0.0 g
Add an additional 1.0 g of sodium chloride	0.2 g	0.8 g
Determine the total amount that dissolves	36.2 g	

- The amount of a substance that dissolves in a given quantity of solvent at a constant temperature is called the substance's solubility at that temperature.
- If a solution contains the *maximum* amount of solute for a given quantity of solvent at a constant temperature, it is called a(n) saturated solution.

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**Liying Dong**



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