

Organic and Inorganic Reactivity

Lecture 1

In the next four lectures I and Dr. Welton will be looking at the reactivity of organic and inorganic compounds respectively in an overall perspective. In particular we want to show you where the patterns of reactivity in these two areas of chemistry are similar and where they differ so that you will be able to see that, in terms of reactivity, organic and inorganic chemistry are related in spite of an historical separation.

Functional Group Approach.

Historically, the chemistry of organic compounds has been taught according to the type of functional group within the molecule. What is a functional group? If we look at 1-butanol:



We see that it has many atoms and therefore potentially many sites of attack by various reagents. However, most (but not all!!) of its chemistry is dictated by the hydroxyl group OH. Since this functions as its main site of reactivity it is termed the functional group. The advantage of teaching according to the functional group is that the chemistry of the functional group is largely (but not exclusively!!) the same irrespective of the nature of the rest of the molecule i.e. 1-butanol behaves just like ethanol, 1- or 2-propanol, 2-butanol, 1- or 2- or 3-pentanol etc. Thus we only have to learn the limited number of reactions of the functional group rather than the individual chemistries of each molecule containing that group. As a means of classification of organic compounds, the functional group approach is still very useful. However, if we look at a list of common functional groups in organic chemistry (Table 1) we see that there is still a lot of bulk learning, especially since this list is not comprehensive and there are at least 100 functional groups.

Table 1 Common Functional Groups

<u>Structure</u>	<u>Name</u>
R-F	Fluoroalkanes, Alkyl fluorides
R-Cl	Chloroalkanes, Alkyl chlorides
R-Br	Bromoalkanes, Alkyl bromides
R-I	Iodoalkanes, Alkyl iodides
R-OH	Alkanols, Alcohols
R ¹ -OR ²	Dialkyl ethers, Ethers
R-SH	Alkanethiols, Thiols
R ¹ -SR ²	Dialkyl sulphides
RSO ₃ H	Sulphonic Acids
R-NH ₂	Monoalkylamines, Primary amines
R ₂ NH	Dialkylamines, Secondary amines
R ₃ N	Trialkylamines, Tertiary amines
R ₄ N ⁺ X ⁻	Tetraalkylammonium Salts

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**University of Illinois (Urbana-
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